



CHEMISTRY

CHEAT SHEET



LEWIS STRUCTURES

LEWIS STRUCTURES CHEAT SHEET

Lewis Structures are simplified representations of the valence electrons in molecules or compounds. They are important in understanding molecular structure and bonding. Follow the steps below to learn how to draw them.

The periodic table highlights the following groups for Lewis structures:

- Group 1: Alkali metals
- Group 2: Alkaline earth metals
- Groups 13-18: Main group elements
- Group 18: Noble gases

Directions

1. Identify the central atom in the molecule or compound.
2. Count the total number of valence electrons in the molecule or compound.
3. Place the central atom in the center.
4. Place the surrounding atoms around the central atom.
5. Connect the atoms with single bonds.
6. Distribute the remaining electrons as lone pairs on the outer atoms.
7. If the central atom has a lone pair, place it on the central atom.
8. Check the octet rule for all atoms.

Exceptions

1. Hydrogen: Always has 2 valence electrons.
2. Boron: Can have 3 valence electrons.
3. Oxygen: Can have 6 valence electrons.
4. Nitrogen: Can have 5 valence electrons.
5. Halogens: Can have 7 valence electrons.

Formulas for Lewis Structures with 1-8 Valence Electrons

1 Valence Electrons		2 Valence Electrons	
3 Valence Electrons		4 Valence Electrons	
5 Valence Electrons		6 Valence Electrons	
7 Valence Electrons		8 Valence Electrons	

Examples: Draw the Lewis Structures for Magnesium (Mg), Boron (B), and Neon (Ne)

Magnesium has 2 valence electrons



Boron has 3 valence electrons



Neon has 8 valence electrons



Remember: Always check the octet rule for all atoms. Some elements can have more than 8 valence electrons.

Lewis Structure Chemistry Study Guide

Chao Zhang



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