

Guided Practice: Equilibrium

Collision Theory & Reaction Rates

A - How Chemical Reactions Occur

1. Using the collision model, explain at a molecular level why reactions proceed FASTER when concentrations increase.

B - Conditions that Affect Reaction Rates

2. What is activation energy? After consulting your text (Fig 17.2 and 17.3) draw a graph showing a reaction pathway - labeling the activation energy within the graph.



3. Use the collision theory to explain why reactions should occur more SLOWLY at lower temperatures.
4. How does a catalyst speed up a reaction? (Refer to the activation energy labeled in #2 in your answer.)
5. Explain the difference between a homogeneous reaction and a heterogeneous reaction. Provide an example of each.
6. List the factors that affect reaction rates; explain how each factor can affect the rate of a reaction:
- a.
 - b.
 - c.
 - d.
7. After reading the definition for equilibrium, provide an example of equilibrium in everyday life.
8. Using your understanding of equilibrium, provide your own definition for chemical equilibrium.

Equilibrium Unit Webquest And Notes

L Towne



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