

HOLT MODERN CHEMISTRY -- CHAPTER 2 REVIEW

The following pages contain the bulk (but not all) of the information for the chapter 2 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and **NOTHING** else, you should at least be able to **PASS** the test. ***** Test items will be recall, examples, and/or application of this content. *****

OUTCOMES

- 2.1: Describe the scientific method process. *(T & R)*
- 2.2: Measure length, volume, mass, and time using metric units. ($1 \text{ mL} = 1 \text{ cm}^3$ & $1 \text{ L} = 1 \text{ dm}^3$) *(T & R)*
- 2.2: Use volume and mass to determine density; compare relative densities of solids, liquids, gases; compare relative densities of pure liquids and solutions. *(F & PK)*
- 2.2: Use density to identify substances in activities involving collection of data and precision and accuracy in measurements. *(F & PK)*
- 2.3: Distinguish between precision and accuracy. *(T & R)*
- 2.3: Determine and perform calculations using significant figures. *(F & PK)*
- 2.3: Convert from numbers to scientific notation and reverse. *(F & PK)*

2.1 SCIENTIFIC METHOD

- **Vocabulary**
 - **scientific method** -- a series of steps followed to solve problems, including collecting data, formulating a hypothesis, testing the hypothesis, and stating conclusions
 - **system** -- a set of particles or interacting components considered to be a distinct physical entity for the purpose of study
 - **hypothesis** -- an explanation that is based on prior scientific research or observations and that can be tested
 - **model** -- a pattern, plan, representation, or description designed to show the structure or workings of an object, system, or concept
 - **theory** -- an explanation for some phenomenon that is based on observation, experimentation, and reasoning
- **Chapter Highlights**
 - The scientific method is a logical approach to solving problems that lend themselves to investigation.
 - A hypothesis is a testable statement that serves as the basis for predictions and further experiments. It is usually written as an if/then statement.
 - A theory is a broad generalization that explains a body of known facts or phenomena.
 - Communication is a very important aspect of all parts of the scientific method, including successful and/or unsuccessful results.

2.2 UNITS OF MEASUREMENT

- **Vocabulary**
 - **quantity** -- something that has magnitude, size, or amount
 - **SI** -- Le Systeme International d'Unites, or the International System of Units, which is the measurement system that is accepted worldwide
 - **weight** -- a measure of the gravitational force exerted on an object; its value can change with the location of the object in the universe
 - **derived unit** -- a unit of measure that is a combination of other measurements
 - **volume** -- a measure of the size of a body or region in three-dimensional space
 - **density** -- the ratio of the mass of a substance to the volume of the substance; often expressed as grams per cubic centimeter for solids and liquids and as grams per liter for gases
 - **conversion factor** -- a ratio that is derived from the equality of two different units and that can be used to convert from one unit to the other
 - **dimensional analysis** -- a mathematical technique for studying dimensions of physical quantities

Holt Chemistry Chapter 2 Test

Darrell D. Ebbing



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Holt Chemistry Ralph Thomas Myers, 2004 **KY HS Test Prac Wkbks W/Corr Sci 2001** Holt Rinehart & Winston, 2002-03 *Principles of Modern Chemistry* David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **Laboratory Experiments to Accompany "Modern Chemistry,"** Fredus Nelson Peters, 1902

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concepts and key problem solving skills through critical thinking Math and problem solving are covered early in the text Corwin builds your problem solving ability through innovative learning aids and technology formulated to meet your needs This revision retains all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking Handbook of Modern Chemistry Inorganic and Organic for the Use of Students Charles Meymott Tidy,1887 **Lab Experiments Modern Chemistry** Harold Clark Metcalfe,1982 *Addison-Wesley Chemistry* ,1990 **A Diagnostic Study of the Subject Matter of High School Chemistry** Samuel Ralph Powers,1924 *Lab Experiments for Modern Chemistry* Tzimopoulo,1990 **Books in Print Supplement** ,2002 **Essentials of Modern Chemistry** Charles Elwood Dull,1918 *Study Guide* Joan I. Senyk,Darrell D. Ebbing,1984 *Modern Chemistry* H. Clark Metcalfe,Joseph F. Castka,1966 *General Chemistry Study Guide Sixth Edition* Darrell D. Ebbing,1999 **Children's Books in Print, 2007** ,2006 **The Nature of the Chemical Concept** Keith S Taber,2022-06-29 This book offers a step by step analysis and discussion of just why some students find chemistry difficult by examining the nature of chemistry concepts and how they are communicated and learnt

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