



# Covalent Compounds Unit Study Guide

Name \_\_\_\_\_

Teacher \_\_\_\_\_

Date: \_\_\_\_\_

Period \_\_\_\_\_

**Reference:**

Chemistry in the Community, pp 181-183,  
Notes  
Naming covalent compounds worksheet  
Viscosity Lab  
Exploring Properties of Ionic and Covalent compounds

**Vocabulary**

**Understand and apply the definitions of the following terms:**

Covalent

Valence electrons

e dot diagrams (Lewis dot diagrams)

Polar

Nonpolar

Intermolecular forces

Octet rule

Polymers

Viscosity

Boiling points

**Targeted Skills**

- Diagram a covalent molecule using a e dot diagrams (Lewis dot diagrams).
- Know the following properties:
  - > Ionic compounds = the transfer of electron(s), i.e. "gain or lose"  
Metal cations & nonmetal anion (or polyatomic ion)  
Net charge = zero (criss cross charges)  $X^{+n}Y^{-m} = X_nY_m$   
Dissociate (dissolve) in water
  - > Molecular compounds = sharing electrons by forming a covalent bond  
2 or more nonmetals  
Prefixes indicate # of each atom  
Polar molecules have a  $\delta^+$  & a  $\delta^-$   
Polar molecules dissolve in water  
Nonpolar molecules do not have oppositely charged ends, therefore do not dissolve in water.
  - > Octet rule = atoms share, gain, or lose e- to obtain full outer energy level (usually 8 e-)
- Name covalent compounds using prefixes: (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca).
- Compare/contrast the arrangement of atoms in molecules to the arrangement of atoms in ionic compounds.
- Describe the influence of intermolecular forces on physical and chemical properties of covalent compounds. Example: Boiling points, viscosity, and density.
- Be able to describe properties (boiling points, melting points, viscosity, solubility, conductivity) of polymers, ionic crystals, metallic substances, atoms and molecules.

**Make sure you remember from past tests?**

Atomic structure

Significant figures

Accuracy &amp; precision

Mixtures

Nomenclature

Solubility

Safety

Elements

Ionic compound formula writing

# Holt Chemistry Covalent Compounds Study Guide

**Jessica J Manson**



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