

### Homework 1

<p>The average human lung expands by about 0.50 LL during each breath.</p>	<p>Work = <math>-p \times \Delta</math>  <math>= -(1.0) \times (0.50L) \text{atm}</math>  <math>= -0.50L \times \text{atm}</math>                      Conversion of <math>1 L \times \text{atm} \rightarrow J = /101.325</math>  <math>= -50.66</math> or rounded to <math>-51</math></p>
<p>The air within a piston equipped with a cylinder absorbs 545 J of heat and expands from an initial volume of 0.11 LL to a final volume of 0.88 LL against an external pressure of 1.0 barbar.</p>	<p><math>W = -p \times \Delta V</math>      <math>\Delta V = 0.88 - 0.11 = .77L</math>  <math>= -1 \times 0.77</math>  <math>= -77L \times \text{atm} \times 101.325 = -78.02J</math></p> <p><math>\Delta E =</math> change in internal energy    <math>q = \text{heat} = 545J</math>  <math>w = 78.02</math>  <math>\Delta E = q + w</math>  <math>= 545 + (-78.02)</math>  <math>= 466.98J</math>  <math>= 467J</math></p>
<p>A gas is compressed from an initial volume of 5.35 LL to a final volume of 1.23 LL by an external pressure of 1.00 barbar . During the compression, the gas releases 121 J of heat.</p>	<p><math>\Delta E = q + w</math>      <math>q = -121</math>                      Calc work done <math>w = -p \times \Delta V</math>  <math>1 \text{ atm} = 1.01 \times 10^5 \text{ Pa}</math>, and <math>1 \text{ liter} = 10^{-3} \text{ m}^3</math>, so:  <math>W = (1.01 \times 10^5 \text{ N/m}^2) \times (5.35 - 1.24) \times 10^{-3} \text{ m}^3</math>  <math>= 415.11J</math></p> <p>So <math>\Delta E = -121 + 415.11</math>  <math>294</math> But with correct values its <math>291</math></p>
<p>What is the work associated with decomposition of trinitrotoluene (TNT) upon detonation according to the following reaction at 281.65 KK?  <math>C_7H_5N_3O_6(s) \rightarrow 3N_2(g) + 5H_2(g) + 3CO_2(g) + 7C(s)</math></p>	<p>1 mole <math>C_7H_5N_3O_6</math> gives <math>3/2</math> moles <math>N_2</math> and <math>5/2</math> moles <math>H_2</math> and 3 moles <math>O_2</math> gas                      Total gas moles produced = <math>N_2</math> moles + <math>H_2</math> moles + <math>O_2</math> moles  <math>= 1.5 + 2.5 + 3 = 7</math></p> <p>work done = <math>-P \times dV</math> where <math>dV</math> is change in volume                      we use ideal gas equation <math>PV = nRT</math> ( since <math>V</math> is nothing but volume occup all gases)                      work done = <math>-nRT</math> where <math>R =</math> gas constant = <math>0.08206 \text{ liter atm/molK}</math>  <math>= -7 \text{ mol} \times 0.08206 \text{ liter atm/molK} \times 281.65</math>  <math>= -162.8500 \text{ atm liter /mol}</math>                      we have <math>1 \text{ atm liter} = 101.325 J</math>                      hence work done = <math>-162.8500 \text{ atm liter /mol} \times (101.325 J/\text{atmliter})</math>  <math>= -16500 J/\text{mol}</math>  <math>= -16.500 \text{ KJ/mol}</math>                      ( -ve sign indicates expansion work )                      (correct answer is <math>-16.39</math>)</p>

# Mastering Chemistry Answers Assignment 1

**A Gutmann**



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