

## Part A

Calculate the amount of heat (in kilojoules) required to vaporize 2.58 kg of water at its boiling point. Express the heat in kilojoules to three significant figures.

$$q = 5830 \text{ kJ}$$

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### Correct

At its boiling point, the enthalpy of vaporization ( $\Delta H_{\text{vap}}$ ) of water is 40.7 kJ/mol. Once the given mass,  $2.58 \times 10^3$  g, is converted to the number of moles using the molar mass of water (18.02 g/mol), multiplying the number of moles by  $\Delta H_{\text{vap}}$  yields the amount of heat required to evaporate that amount of water.

## Part B

Suppose that 0.48 g of water at 25°C condenses on the surface of a 55-g block of aluminum that is initially at 25°C. If the heat released during condensation goes only toward heating the metal, what is the final temperature (in degrees Celsius) of the metal block? (The specific heat capacity of aluminum,  $C_{s,\text{Al}}$ , is 0.903 J/(g · °C).)

Express the temperature in degrees Celsius to two significant figures.

$$T_f =$$

°C

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# Mastering Chemistry Answers Chapter 12

**Huangqi Zhang**



## Mastering Chemistry Answers Chapter 12:

## Unveiling the Magic of Words: A Overview of "**Mastering Chemistry Answers Chapter 12**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their power to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Mastering Chemistry Answers Chapter 12**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound affect the souls of its readers.

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