

SOLUTION

IDENTIFY and SET UP: This problem involves the properties of the Carnot cycle and those of an ideal gas. We are given the number of moles n and the pressure and temperature at point a (which is at the higher of the two reservoir temperatures); we can find the volume at a using the ideal-gas equation $pV = nRT$. We then find the pressure and volume at points b , c , and d from the known doubling of volume in step $a \rightarrow b$, from equations given in this section, and from $pV = nRT$. In each step we use Eqs. (20.10) and (20.11) to find the heat flow and work done and Eq. (19.13) to find the internal energy change.

EXECUTE: The volume at point a is

$$V_a = \frac{nRT_a}{p_a} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K})}{(1.00 \times 10^5 \text{ Pa})} = 8.31 \times 10^{-3} \text{ m}^3$$
 The volume at point b is $V_b = 2V_a = 16.6 \times 10^{-3} \text{ m}^3$. The pressure at point b is

$$p_b = \frac{nRT_b}{V_b} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K})}{16.6 \times 10^{-3} \text{ m}^3} = 5.00 \times 10^4 \text{ Pa}$$
 The volume at point c is $V_c = 2V_b = 33.2 \times 10^{-3} \text{ m}^3$. The pressure at point c is

$$p_c = \frac{nRT_c}{V_c} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(300 \text{ K})}{33.2 \times 10^{-3} \text{ m}^3} = 1.50 \times 10^4 \text{ Pa}$$
 The volume at point d is $V_d = 2V_c = 66.4 \times 10^{-3} \text{ m}^3$. The pressure at point d is

$$p_d = \frac{nRT_d}{V_d} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(300 \text{ K})}{66.4 \times 10^{-3} \text{ m}^3} = 7.50 \times 10^3 \text{ Pa}$$

The heat flow Q_{ab} and work done W_{ab} in the isothermal expansion $a \rightarrow b$ are

$$Q_{ab} = nRT_a \ln \frac{V_b}{V_a} = (0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K}) \ln 2 = 582 \text{ J}$$

$$W_{ab} = nRT_a \ln \frac{V_b}{V_a} = 582 \text{ J}$$
 The heat flow Q_{bc} and work done W_{bc} in the adiabatic expansion $b \rightarrow c$ are

$$Q_{bc} = 0$$

$$W_{bc} = -\Delta U_{bc} = -nC_V \Delta T = -(0.200 \text{ mol})(12.47 \text{ J/mol}\cdot\text{K})(300 \text{ K} - 500 \text{ K}) = 503 \text{ J}$$

The heat flow Q_{cd} and work done W_{cd} in the isothermal expansion $c \rightarrow d$ are

$$Q_{cd} = nRT_c \ln \frac{V_d}{V_c} = (0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(300 \text{ K}) \ln 2 = 351 \text{ J}$$

$$W_{cd} = nRT_c \ln \frac{V_d}{V_c} = 351 \text{ J}$$
 The heat flow Q_{da} and work done W_{da} in the adiabatic compression $d \rightarrow a$ are

$$Q_{da} = 0$$

$$W_{da} = -\Delta U_{da} = -nC_V \Delta T = -(0.200 \text{ mol})(12.47 \text{ J/mol}\cdot\text{K})(500 \text{ K} - 300 \text{ K}) = -503 \text{ J}$$

The heat flow Q_{ab} and work done W_{ab} in the isothermal expansion $a \rightarrow b$ are

$$Q_{ab} = nRT_a \ln \frac{V_b}{V_a} = (0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K}) \ln 2 = 582 \text{ J}$$

$$W_{ab} = nRT_a \ln \frac{V_b}{V_a} = 582 \text{ J}$$

For the adiabatic expansion $b \rightarrow c$, $Q_{bc} = 0$. From the first law of thermodynamics, $\Delta U_{bc} = Q_{bc} - W_{bc} = -W_{bc}$; the work W_{bc} done by the gas in this adiabatic expansion equals the negative of the change in internal energy of the gas. From Eq. (19.13) we have $\Delta U = nC_V \Delta T$, where $\Delta T = T_c - T_b$. Using $C_V = 20.8 \text{ J/mol}\cdot\text{K}$ for an ideal diatomic gas, we find

$$W_{bc} = -\Delta U_{bc} = -nC_V(T_c - T_b) = nC_V(T_b - T_c) \\ = (0.200 \text{ mol})(20.8 \text{ J/mol}\cdot\text{K})(500 \text{ K} - 300 \text{ K}) = 832 \text{ J}$$

For the isothermal compression $c \rightarrow d$, $\Delta U_{cd} = 0$; Eq. (20.11)

gives

$$Q_{cd} = -W_{cd} = nRT_c \ln \frac{V_d}{V_c} = (0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(300 \text{ K}) \ln 2 = 351 \text{ J}$$
 The heat flow Q_{da} and work done W_{da} in the adiabatic compression $d \rightarrow a$ are

$$Q_{da} = 0$$

$$W_{da} = -\Delta U_{da} = -nC_V \Delta T = -(0.200 \text{ mol})(20.8 \text{ J/mol}\cdot\text{K})(500 \text{ K} - 300 \text{ K}) = -832 \text{ J}$$

Step	Process	Heat Flow Q (J)	Work Done W (J)
1	$a \rightarrow b$	582	582
2	$b \rightarrow c$	0	832
3	$c \rightarrow d$	351	351
4	$d \rightarrow a$	0	-832

The heat flow Q_{ab} and work done W_{ab} in the isothermal expansion $a \rightarrow b$ are

$$Q_{ab} = nRT_a \ln \frac{V_b}{V_a} = (0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K}) \ln 2 = 582 \text{ J}$$

$$W_{ab} = nRT_a \ln \frac{V_b}{V_a} = 582 \text{ J}$$

IDENTIFY and SET UP: The Carnot cycle is a cycle that operates between two reservoirs at different temperatures. The cycle consists of four processes: two isotherms and two adiabats. The heat flow Q_{ab} and work done W_{ab} in the isothermal expansion $a \rightarrow b$ are

$$Q_{ab} = nRT_a \ln \frac{V_b}{V_a} = (0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K}) \ln 2 = 582 \text{ J}$$

$$W_{ab} = nRT_a \ln \frac{V_b}{V_a} = 582 \text{ J}$$

Now Work PROBLEM 19

PROBLEM 19 A Carnot cycle is shown in the pV diagram in Figure 20.19. The cycle consists of four processes: two isotherms and two adiabats. The heat flow Q_{ab} and work done W_{ab} in the isothermal expansion $a \rightarrow b$ are

$$Q_{ab} = nRT_a \ln \frac{V_b}{V_a} = (0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K}) \ln 2 = 582 \text{ J}$$

$$W_{ab} = nRT_a \ln \frac{V_b}{V_a} = 582 \text{ J}$$

$$Q_{bc} = 0$$

$$W_{bc} = -\Delta U_{bc} = -nC_V \Delta T = -(0.200 \text{ mol})(12.47 \text{ J/mol}\cdot\text{K})(300 \text{ K} - 500 \text{ K}) = 503 \text{ J}$$

Masteringphysics Solution Chapter

David Reid, James Walker



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Essential University Physics Richard Wolfson, 2007 Richard Wolfson's *Essential University Physics* is a concise and progressive calculus based physics textbook that offers clear writing, great problems and interesting real life applications. At nearly half the length and half the price of other physics texts on the market, *Essential University Physics* is a compelling alternative for professors who want to focus on the fundamentals. *Doing Physics 1*: Mechanics, Motion in a Straight Line, Motion in Two and Three Dimensions, Force and Motion Using Newton's Laws, Work, Energy and Power, Conservation of Energy, Gravity, Systems of Particles, Rotational Motion, Rotational Vectors and Angular Momentum, Static Equilibrium. Part 2: Oscillations, Waves and Fluids, Oscillatory Motion, Wave Motion, Fluid Motion, Thermodynamics, Temperature and Heat, The Thermal Behavior of Matter, Heat Work and the First Law of Thermodynamics, The Second Law of Thermodynamics. For all readers interested in calculus based physics.

College Physics Randall Dewey Knight, Brian Jones, Stuart Field, 2007

Proceedings of the Blended Learning in Science, Teaching and Learning Symposium, 2005. Presents proceedings of the annual Uniserve Conference. The papers contained in this book include topics as teaching science online, tutorial benefits of online assignments, blended learning and other related issues in relation to teaching science at a university level.

Principles of Physics Gordon Lind, Raymond A. Serway, 1997-10. The second edition like the first follows the guidelines of the Introductory University Physics Project IUPP. The revision includes a stronger conceptual approach offering new conceptual examples and problems and it presents contemporary physics topics early to gain student interest. This book is intended for the science and engineering physics course.

[Pocket Guide to Accompany College Physics, Fourth Edition, by Serway & Faughn](#) V. Gordon Lind, 1994

College Physics Raymond A. Serway, V. Gordon Lind, 1998-08. This 5 by 7 paperback is a section by section capsule of the textbook that provides a handy guide for looking up important concepts, equations and problem solving hints.

Pocket Guide to Accompany Principles of Physics Raymond A. Serway, V. Gordon Lind, 1994

High School Physics Oswald Hance Blackwood, 1958

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 - Chapter 4: Masteringphysics Solution Chapter in Specific Contexts
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