

Name _____

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Modern Chemistry • CHAPTER 7

HOMework 7-1

(pp. 261–266)

VOCABULARY

Write *true* or *false* for each statement.

1. A **chemical formula** indicates the relative number of molecules of each kind in a chemical compound.

2. **Monoatomic ions** are ions formed from a single atom. _____
3. **Monoatomic anions** of an element are identified by the element's name. _____
4. A **binary compound** is made of two or more different elements. _____
5. The **Stock system** is a system for naming chemical elements and their ions. _____

SKILL BUILDER

Explain what each chemical formula tells you about the composition of the compound. The first one is done for you.



The subscript 3 refers to the number of oxygen atoms. There are one atom of sodium and one atom of nitrogen for every three atoms of oxygen.



STANDARDIZED TEST PREP

Circle the letter of the best answer.

1. Which statement is true?
 - a. All main-group elements are able to form ions.
 - b. The total numbers of positive charges and negative charges must be equal in a binary compound.
 - c. The nonmetals of groups 15, 16, and 17 lose electrons to form cations.
 - d. Elements in the *d* block form ions of only one charge.
2. Why do many main-group elements lose or gain electrons to form ions?
 - a. The elements are unstable and therefore lose electrons.
 - b. Their tendency is to maximize potential energy.
 - c. The elements form covalent bonds, sharing their electrons with other atoms.
 - d. Their tendency is to form a complete outermost octet in a noble-gas configuration.

Modern Chemistry Chapter 7 Homework Answers

Patrick Vollmar



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