

Chapter 5: Gases - Things to know*

Please remember that simply knowing or memorizing the information below is not sufficient. You must be able to know when and how to APPLY the items below appropriately.

5.1 Basics

Pressure and Units *(see guide!)*

5.2 Simple Gas Laws *(see guide!)*

5.3 Ideal Gas Law *(see preview below!)*

Used to determine one gas variable for a given scenario, knowing the other variables.

A. $PV=nRT$ ($R = 0.0821 \text{ L atm/mol K}$; $8.314 \text{ L kPa/mol K}$; be sure to match all units to the constant)

B. Limitations:

1. Works well at low pressures *as the molecules of the gas rarely collide or come into close enough proximity for intermolecular forces to be significant*
2. Works well at high temperatures *as the molecules of the gas have so much energy that the intermolecular forces and energy lost in collisions is negligible*
3. Most gases do not behave ideally above 1 atm pressure
4. Does not work well near the condensation conditions of a gas
5. Deviations occur due to stronger attractive forces

C. Derivation to relate molar mass and density

$D = m/V$ $n = m/M$ where M is molar mass

$PV = m/MRT$; $M = DRT/P$

5.4 Gas Mixtures and Collecting Gas of Water

5.5 Diffusion and Effusion

Modern Chemistry Gases Study Guide

WJ Hussar

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