



CHEMISTRY CHEAT SHEET



EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c .

What is Equilibrium?
 A reversible reaction is one that can proceed in either direction. At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction. The concentrations of the reactants and products remain constant over time.

Equilibrium Constant K_c
 The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of their stoichiometric coefficients. It is a constant for a given reaction at a given temperature.



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Reaction	Equilibrium Constant K_c
$aA + bB \rightleftharpoons cC + dD$	$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
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Factor	Effect on Equilibrium	Effect on K_c
Concentration	<ul style="list-style-type: none"> Increasing the concentration of a reactant or product will shift the equilibrium position to favor the formation of the other. 	<ul style="list-style-type: none"> K_c is not affected by changes in concentration.
Pressure	<ul style="list-style-type: none"> Increasing the pressure will shift the equilibrium position to favor the formation of the side with fewer moles of gas. 	<ul style="list-style-type: none"> K_c is not affected by changes in pressure.
Temperature	<ul style="list-style-type: none"> Increasing the temperature will shift the equilibrium position to favor the endothermic reaction. 	<ul style="list-style-type: none"> K_c is affected by changes in temperature.

Modern Chemistry Study Guide Chemical Equilibrium

J Dewey



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