

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



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Table of Contents Modern Chemistry Study Guide Mixed Review Answers

1. Understanding the eBook Modern Chemistry Study Guide Mixed Review Answers
 - The Rise of Digital Reading Modern Chemistry Study Guide Mixed Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Study Guide Mixed Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Study Guide Mixed Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Study Guide Mixed Review Answers
 - Personalized Recommendations
 - Modern Chemistry Study Guide Mixed Review Answers User Reviews and Ratings

- Modern Chemistry Study Guide Mixed Review Answers and Bestseller Lists
- 5. Accessing Modern Chemistry Study Guide Mixed Review Answers Free and Paid eBooks
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 - Interactive Elements Modern Chemistry Study Guide Mixed Review Answers
- 8. Staying Engaged with Modern Chemistry Study Guide Mixed Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
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- 9. Balancing eBooks and Physical Books Modern Chemistry Study Guide Mixed Review Answers
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- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modern Chemistry Study Guide Mixed Review Answers
 - Setting Reading Goals Modern Chemistry Study Guide Mixed Review Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modern Chemistry Study Guide Mixed Review Answers
 - Fact-Checking eBook Content of Modern Chemistry Study Guide Mixed Review Answers
 - Distinguishing Credible Sources

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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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