

**CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY**1) Which of the following are weak electrolytes? HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ , KCl

A) HCl, KCl

B) HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ , KClC)  $\text{HC}_2\text{H}_3\text{O}_2$ , KClD)  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ E) HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ , KCl

Answer: D Diff: 2

2) What are the spectator ions in the reaction between KOH (aq) and  $\text{HNO}_3$ (aq)?A)  $\text{K}^+$  and  $\text{H}^+$ C)  $\text{K}^+$  and  $\text{NO}_3^-$ E)  $\text{OH}^-$  onlyB)  $\text{H}^+$  and  $\text{OH}^-$ D)  $\text{H}^+$  and  $\text{NO}_3^-$ 

Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of  $\text{CaCO}_3$  when aqueous solutions of  $\text{Na}_2\text{CO}_3$  and  $\text{CaCl}_2$  are mixed is \_\_\_\_\_.A)  $2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Na}_2\text{CO}_3(\text{aq})$ 

E)

B)  $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq})$  $\text{Na}_2\text{CO}_3(\text{aq}) + \text{CaCl}_2(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{CaCO}_3(\text{s})$ C)  $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{NaCl}(\text{aq})$ D)  $\text{Ca}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CaCO}_3(\text{s})$ 

Answer: D Diff: 3

4) Which of the following are strong acids? HI,  $\text{HNO}_3$ , HF, HBr

A) HF, HBr

C) HI, HF, HBr

E) HI,  $\text{HNO}_3$ , HBrB) HI,  $\text{HNO}_3$ , HF, HBrD)  $\text{HNO}_3$ , HF, HBr

Answer: E Diff: 2

5) A neutralization reaction between an acid and a metal hydroxide produces \_\_\_\_\_.

A) water and a salt

C) oxygen gas

E) ammonia

B) hydrogen gas

D) sodium hydroxide

Answer: A Diff: 2

6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na

B) Au

C) Fe

D) Ca

E) Ag

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?



A) Ni

B) Ag

C) Cu

D) Pb

E) N

Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100

B) 0.0500

C) 0.0333

D) 0.0250

E) 125

Answer: A Diff: 3

# Modern Chemistry Test Answer Key Ch8

**Michael Seilmaier**



## **Modern Chemistry Test Answer Key Ch8:**

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**Modern Electrochemistry 2B** John O'M. Bockris, Amulya K.N. Reddy, 2007-05-08. This book had its nucleus in some lectures given by one of us (J O M B) in a course on electrochemistry to students of energy conversion at the University of Pennsylvania. It was there that he met a number of people trained in chemistry, physics, biology, metallurgy, and materials science, all of whom wanted to know something about electrochemistry. The concept of writing a book about electrochemistry, which could be understood by people with very varied backgrounds, was thereby engendered. The lectures were recorded and written up by Dr. Klaus Müller as a 293-page manuscript. At a later stage, A. K. N. R. joined the effort; it was decided to make a fresh start and to write a much more comprehensive text. Of methods for direct energy conversion, the electrochemical one is the most advanced and seems the most likely to become of considerable practical importance. Thus, conversion to electrochemically powered transportation systems appears to be an important step by means of which the difficulties of air pollution and the effects of an increasing concentration in the atmosphere of carbon dioxide may be met. Corrosion is recognized as having an electrochemical basis. The synthesis of nylon now contains an important electrochemical stage. Some central biological mechanisms have been shown to take place by means of electrochemical reactions. A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States.

Modern Chemistry Holt, Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

Children's Books in Print R. R. Bowker Publishing, Bowker, 1999-12      **Books in Print Supplement**, 2002      Modern Inorganic Chemistry William L. Jolly, 1991      **Modern Chemical Technology** Chemical Technician Curriculum Project. Writing Team, 1972      Customized Modern Inorganic Chemistry William L. Jolly, 1998      **Linne & Ringsrud's Clinical Laboratory Science E-Book** Mary Louise Turgeon, 2018-12-22. Thoroughly updated and easy to follow. Linne Ringsrud's

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*Voet's Principles of Biochemistry, Global Edition* Donald Voet, Judith G. Voet, Charlotte W. Pratt, 2018-01-11. Voet's Principles of Biochemistry Global Edition addresses the enormous advances in biochemistry, particularly in the areas of structural biology and bioinformatics. It provides a solid biochemical foundation that is rooted in chemistry to prepare students for the scientific challenges of the future. New information related to advances in biochemistry and experimental approaches for studying complex systems are introduced. Notes on a variety of human diseases and pharmacological effectors have been expanded to reflect recent research findings. While continuing in its tradition of presenting complete and balanced coverage, this Global Edition includes new pedagogy and enhanced visuals that provide a clear pathway for student learning.

**Novel Nanomaterials for Biomedical, Environmental and Energy Applications** Xiaoru Wang, Xi Chen, 2018-11-16. Novel Nanomaterials for Biomedical, Environmental and Energy Applications is a comprehensive study on the cutting-edge progress in the synthesis and characterization of novel nanomaterials and their subsequent advances and uses in biomedical

environmental and energy applications Covering novel concepts and key points of interest this book explores the frontier applications of nanomaterials Chapters discuss the overall progress of novel nanomaterial applications in the biomedical environmental and energy fields introduce the synthesis characterization properties and applications of novel nanomaterials discuss biomedical applications and cover the electrocatalytical and photothermal effects of novel nanomaterials for efficient energy applications The book will be invaluable to academic researchers and biomedical clinicians working with nanomaterials Offers comprehensive details on novel and emerging nanomaterials Presents a comprehensive view of new and emerging tactics for the synthesis of efficient nanomaterials Describes and monitors the functions of applications of new and emerging nanomaterials in the biomedical environmental and energy fields

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**School Chemistry Teachers Magazine** ,1967 **Modern Chemistry** Holt, Rinehart and Winston Staff,2006-01

**Sourcebook for Chemistry and Physics** David R. Hittle,1973 **Children's Books in Print, 2007** ,2006 *Modern Instrumental Analysis* Satinder Ahuja,Neil Jespersen,2006 *Modern Instrumental Analysis* covers the fundamentals of instrumentation and provides a thorough review of the applications of this technique in the laboratory It will serve as an educational tool as well as a first reference book for the practicing instrumental analyst The text covers five major sections 1 Overview Sampling Evaluation of Physical Properties and Thermal Analysis2 Spectroscopic Methods 3 Chromatographic Methods 4 Electrophoretic and Electrochemical Methods 5 Combination Methods Unique Detectors and Problem Solving Each section has a group of chapters covering important aspects of the titled subject and each chapter includes applications that illustrate the use of the methods The chapters also include an appropriate set of review questions

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