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Worksheet 8 - Stoichiometry calculations

1. Given the following equation: $2\ C_4H_{10} + 13\ O_2 \rightarrow 8\ CO_2 + 10\ H_2O$, show what the following molar ratios should be.
   a. C_4H_{10} / O_2 — 2/13
   b. O_2 / CO_2 — 13/8
   c. O_2 / H_2O — 13/10
   d. C_4H_{10} / CO_2 — 1/4
   e. C_4H_{10} / H_2O — 1/5

2. Given the following equation: $2\ KClO_3 \rightarrow 2\ KCl + 3\ O_2$

How many moles of O_2 can be produced by letting 12.00 moles of $KClO_3$ react?

$$12.00\ mol_{KClO_3} \cdot \frac{3\ mol_{O_2}}{2\ mol_{KClO_3}} = 18\ mol_{O_2}$$

3. Given the following equation: $2\ K + Cl_2 \rightarrow 2\ KCl$

(a) How many grams of KCl is produced from 2.50 g of K and excess Cl_2? (b) From 1.00 g of Cl_2 and excess K?

(a) $$2.50\ g_K \cdot \frac{1\ mol_K}{39.1\ g_K} \cdot \frac{2\ mol_{KCl}}{2\ mol_K} \cdot \frac{74.5\ g_{KCl}}{1\ mol_{KCl}} = 4.76\ g_{KCl}$$

(b) $$1.00\ g_{Cl_2} \cdot \frac{1\ mol_{Cl_2}}{70.9\ g_{Cl_2}} \cdot \frac{2\ mol_{KCl}}{1\ mol_{Cl_2}} \cdot \frac{74.5\ g_{KCl}}{1\ mol_{KCl}} = 2.10\ g_{KCl}$$

Module 8 Stoichiometry Answer Key

Lingjun Ying

Module 8 Stoichiometry Answer Key:

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