

Molar Ratio Practice Problems

Following each equation are two requests for molar ratios from the equation.



N_2 to H_2 : NH_3 to H_2 :



O_2 to SO_3 : O_2 to SO_2 :



PCl_3 to Cl_2 : PCl_3 to PCl_5 :



NH_3 to N_2 : H_2O to O_2 :



CO to CO_2 : Fe to CO :

Mole to Mole Practice Problems

Here's the equation to use for all three problems:



1) How many moles of H_2O are produced when 5 moles of oxygen are used?

Remember: *Starting with five moles of oxygen and based on the balanced equation, for every 1 mole of oxygen used, two moles of water are produced.*

$$\underline{\hspace{1cm}} \text{ moles O}_2 \times \frac{2 \text{ mole H}_2\text{O}}{1 \text{ mole O}_2} = \underline{\hspace{1cm}} \text{ moles H}_2\text{O}$$

2) If 3.00 moles of H_2O are produced, how many moles of oxygen must be consumed?

3) If 2.5 moles of H_2O are produced, how many moles of hydrogen gas must be used?

4) Suppose 4.00 grams of H_2 were used? How many grams of water would be produced?

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L Darling-Hammond



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