

Molarity = Moles of solute / Liters of Solution (abbreviation = M)

Molality = Moles of solute / Kg of Solvent (abbreviation = m)

Normality = number of equivalent of solute x Molarity of Solution (abbreviation = N)

Mass Percent = mass of solute / mass of solution

1. How many moles of ethyl alcohol, C_2H_5OH , are present in 65 mL of a 1.5 M solution?

$$\frac{65 \text{ mL } C_2H_5OH}{1000 \text{ mL } C_2H_5OH} \times \frac{1 \text{ L } C_2H_5OH}{1 \text{ L } C_2H_5OH} \times \frac{1.5 \text{ mole } C_2H_5OH}{1 \text{ L } C_2H_5OH} = 0.098 \text{ moles } C_2H_5OH$$

2. How many liters of a 6.0 M solution of acetic acid CH_3COOH , contain 0.0030 moles of acetic acid?

$$\frac{0.0030 \text{ moles } CH_3COOH}{6.0 \text{ moles } CH_3COOH} \times \frac{1 \text{ L solution}}{1 \text{ L solution}} = 5.0 \times 10^{-4} \text{ L solution}$$

3. You want 85 g of KOH. How much of a 3.0 m solution of KOH will provide it?

$$\frac{85 \text{ g KOH}}{56.11 \text{ g KOH}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{1 \text{ kg solvent}}{3.0 \text{ mole KOH}} \times \frac{1000 \text{ g solvent}}{1 \text{ kg solvent}} = 5.0 \times 10^2 \text{ g solvent}$$

$$5.0 \times 10^2 \text{ g solvent} + 85 \text{ g KOH} = 590 \text{ g solution}$$

4. If you dissolve 0.70 moles of HCl in enough water to prepare 250 mL of solution, what is the molarity of the solution you have prepared?

$$M = \frac{0.70 \text{ moles HCl}}{0.250 \text{ L solution}} = 2.8 \text{ M HCl}$$

5. A solution is prepared by adding 2.0 L of 6.0 M HCl to 500 mL of a 9.0 M HCl solution. What is the molarity of the new solution? (Remember, the volumes are additive)

$$\frac{2.0 \text{ L solution}}{1 \text{ L solution}} \times \frac{6.0 \text{ moles HCl}}{1 \text{ L solution}} = 12 \text{ moles HCl}$$

$$\frac{0.500 \text{ L solution}}{1 \text{ L solution}} \times \frac{9.0 \text{ moles HCl}}{1 \text{ L solution}} = 4.5 \text{ moles HCl}$$

$$12 \text{ moles HCl} + 4.5 \text{ moles HCl} = 17 \text{ moles}$$

$$2.0 \text{ L solution} + 0.500 \text{ L solution} = 2.5 \text{ L solution}$$

$$M = \frac{17 \text{ moles}}{2.5 \text{ L solution}} = 6.8 \text{ M HCl}$$

6. Convert the following Molarities to Normalities.

a. 2.5 M HCl = 2.5 N

b. 1.4 M H_2SO_4 = 2.8 N

c. 1.0 M NaOH = 1.0 N

d. 0.5 M $Ca(OH)_2$ = 1 N

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Table of Contents Molarity Answer Key Page68

1. Understanding the eBook Molarity Answer Key Page68
 - The Rise of Digital Reading Molarity Answer Key Page68
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Answer Key Page68
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Answer Key Page68
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Answer Key Page68
 - Personalized Recommendations
 - Molarity Answer Key Page68 User Reviews and Ratings
 - Molarity Answer Key Page68 and Bestseller Lists
5. Accessing Molarity Answer Key Page68 Free and Paid eBooks
 - Molarity Answer Key Page68 Public Domain eBooks
 - Molarity Answer Key Page68 eBook Subscription Services
 - Molarity Answer Key Page68 Budget-Friendly Options

6. Navigating Molarity Answer Key Page68 eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity Answer Key Page68 Compatibility with Devices
 - Molarity Answer Key Page68 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Answer Key Page68
 - Highlighting and Note-Taking Molarity Answer Key Page68
 - Interactive Elements Molarity Answer Key Page68
8. Staying Engaged with Molarity Answer Key Page68
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Answer Key Page68
9. Balancing eBooks and Physical Books Molarity Answer Key Page68
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Answer Key Page68
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molarity Answer Key Page68
 - Setting Reading Goals Molarity Answer Key Page68
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Answer Key Page68
 - Fact-Checking eBook Content of Molarity Answer Key Page68
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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