

Molarity:

- a quantitative description of solution concentration.
- Abbreviated M

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

Problems: Show all work and circle your final answer.

1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of solution are required?

$$4.00 \text{ M} = \frac{\text{moles of solute}}{12.0 \text{ L}} \quad \text{moles of solute} = \boxed{48.0 \text{ mol}}$$

2. How many moles of sucrose are dissolved in 250 mL of solution if the solution concentration is 0.150 M?

$$? \text{ L} = 250 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.25 \text{ L}$$

$$0.150 \text{ M} = \frac{\text{moles of solute}}{0.25 \text{ L}} \quad \text{moles of solute} = \boxed{0.038 \text{ mol}}$$

3. What is the molarity of a solution of HNO₃ that contains 12.6 grams HNO₃ in 1.0 L of solution?

$$? \text{ mol HNO}_3 = 12.6 \text{ g HNO}_3 \times \frac{1 \text{ mol HNO}_3}{63.0 \text{ g HNO}_3} = 0.200 \text{ mol HNO}_3$$

$$M = \frac{0.200 \text{ mol HNO}_3}{1.0 \text{ L}} = \boxed{0.200 \text{ M}}$$

4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?

$$0.700 \text{ M} = \frac{\text{moles of solute}}{0.250 \text{ L}} \quad \text{moles of solute} = 0.175 \text{ mol}$$

$$? \text{ g KNO}_3 = 0.175 \text{ mol KNO}_3 \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} = \boxed{17.7 \text{ g KNO}_3}$$

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Lei Huang



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Table of Contents Molarity Worksheet Answers Page 273

1. Understanding the eBook Molarity Worksheet Answers Page 273
 - The Rise of Digital Reading Molarity Worksheet Answers Page 273
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Worksheet Answers Page 273
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Worksheet Answers Page 273
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Worksheet Answers Page 273
 - Personalized Recommendations
 - Molarity Worksheet Answers Page 273 User Reviews and Ratings
 - Molarity Worksheet Answers Page 273 and Bestseller Lists
5. Accessing Molarity Worksheet Answers Page 273 Free and Paid eBooks

- Molarity Worksheet Answers Page 273 Public Domain eBooks
 - Molarity Worksheet Answers Page 273 eBook Subscription Services
 - Molarity Worksheet Answers Page 273 Budget-Friendly Options
6. Navigating Molarity Worksheet Answers Page 273 eBook Formats
- ePub, PDF, MOBI, and More
 - Molarity Worksheet Answers Page 273 Compatibility with Devices
 - Molarity Worksheet Answers Page 273 Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Molarity Worksheet Answers Page 273
 - Highlighting and Note-Taking Molarity Worksheet Answers Page 273
 - Interactive Elements Molarity Worksheet Answers Page 273
8. Staying Engaged with Molarity Worksheet Answers Page 273
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Worksheet Answers Page 273
9. Balancing eBooks and Physical Books Molarity Worksheet Answers Page 273
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Worksheet Answers Page 273
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molarity Worksheet Answers Page 273
- Setting Reading Goals Molarity Worksheet Answers Page 273
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Worksheet Answers Page 273
- Fact-Checking eBook Content of Molarity Worksheet Answers Page 273
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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