

- B) 27.2 mol
- C) 6.80 mol
- D) 5.44 mol**
- E) 2.27 mol

Determine mole ratio

$$\begin{array}{l} F_2 : NH_3 \\ 5 : 2 \\ 13.6 \text{ mol} : 5.44 \text{ mol} \end{array}$$

4. Potassium chlorate (used in fireworks, flares, and safety matches) forms oxygen and potassium chloride when heated.



How many grams of oxygen are formed when 26.4 g of potassium chlorate is heated?

- A) 223 g
- B) 99.1 g
- C) 10.3 g**
- D) 6.86 g
- E) 4.60 g

Balance the equation



Find number of moles potassium chlorate

$$\text{Molar mass of } KClO_3 = 39.1 + 35.5 + (3 \times 16) = 122.6 \text{ g/mol}$$

$$\text{Moles of } KClO_3 = \frac{26.4 \text{ g}}{122.6 \text{ g/mol}} = 0.2153 \text{ mol}$$

Determine the mole ratio

$$\begin{array}{l} KClO_3 : O_2 \\ 2 : 3 \\ 0.2153 \text{ mol} : 0.323 \text{ mol} \end{array}$$

Find the mass of oxygen formed

$$\text{Molar mass } O_2 = 16 + 16 = 32 \text{ g/mol}$$

$$\text{Mass } O_2 = 32 \text{ g/mol} \times 0.323 \text{ mol} = 10.3 \text{ g}$$

5. Aluminum oxide (used as an adsorbent or a catalyst for organic reactions) forms when aluminum reacts with oxygen.



A mixture of 82.49 g of aluminum (M = 26.98 g/mol) and 117.65 g of oxygen (M = 32.00 g/mol) is allowed to react. What mass of aluminum oxide (M = 101.96 g/mol) can be formed?

- A) 155.8 g**
- B) 200.2 g

Molarity Ws Answers

L Cohen



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