

1 Mole to Mole Relationship between Cu and  
 Ag\_Essay1\_Experimental Data Table  
 Record the following data in the table below.

|  |                  |
|--|------------------|
| mass of empty beaker   | <b>110.000 g</b> |
| mass of beaker with silver nitrate<br>(after all additions) and copper | <b>331.634 g</b> |
| mass of beaker with silver   | <b>113.395 g</b> |

2 Mole to Mole Relationship between Cu and  
 Ag\_Essay2\_Moles Silver Nitrate Added to Reaction  
 How many moles of silver nitrate were added to the reaction?

$$n = M \times V$$

$$n = 0.2 \text{ M} \times 0.200 \text{ L} = 0.04 \text{ mol AgNO}_3$$

3 Mole to Mole Relationship between Cu and  
 Ag\_Essay3\_Silver Ions Consumed

Did all the silver ions get consumed in the reaction? The molar mass of silver is 107.87 g/mol. Justify your answer.

$$n_{\text{Ag}^+} = M \times V$$

$$n_{\text{Ag}^+ \text{ added}} = 0.2 \text{ M} \times 0.200 \text{ L} = 0.04 \text{ mol Ag}^+$$

$$n_{\text{Ag}^+ \text{ formed}} = \frac{m}{MM} = \frac{113.395 \text{ g} - 110.000 \text{ g}}{107.87 \text{ g/mol}} = 0.03147 \text{ mol Ag}$$

**No, not all the silver ions were consumed in the reaction.**

4 Mole to Mole Relationship between Cu and  
 Ag\_Essay4\_Silver Nitrate Plus Zinc Reaction

# Mole Relationship Answers

**Mary Millhollon, Richard H. Langley**

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