

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

- copper (II) oxide CuO
- mercury (I) oxide Hg₂O
- gold (III) chloride AuCl₃
- thallium (III) bromide TlBr₃
- bismuth (V) oxide Bi₂O₅
- terbium (III) oxide Tb₂O₃
- uranium (VI) oxide UO₃
- protactinium (IV) bromide PaBr₄
- cerium (III) oxide Ce₂O₃
- arsenic (V) sulphide As₂S₅

- manganese (III) oxide Mn₂O₃
- vanadium (II) bromide VBr₂
- niobium(V) oxide Nb₂O₅
- titanium (III) oxide Ti₂O₃
- titanium (III) nitride TiN
- iron (II) oxide FeO
- cobalt (II) phosphide Co₃P₂
- tin (II) oxide SnO
- thulium (II) bromide TmBr₂
- copper (II) bromide CuBr₂

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

- SnCl₄ Tin (IV) Chloride
- BiBr₅ Bismuth (V) Bromide
- PoO₂ Polonium (IV) Oxide
- PbI₂ Lead (II) Iodide
- HgO Mercury (II) Oxide
- HgCl Mercury (I) Chloride
- Au₂O₃ Gold (III) Oxide
- FeCl₂ Iron (II) Chloride

- PdF₄ Palladium (IV) Fluoride
- Os₂O₃ Osmium (III) Oxide
- MoBr₂ Molybdenum (II) Bromide
- VCl₅ Vanadium (V) Chloride
- Mn₂O₃ Manganese (III) Oxide
- CoO Cobalt (II) Oxide
- Np₂O₃ Neptunium (III) Oxide
- V₂O₅ Vanadium (V) Oxide

Molecular Geometry Packet Answers

Ronald J Gillespie, Istvan Hargittai



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Electron Dynamics In Molecular Interactions: Principles And Applications Frank Hagelberg,2013-12-23 This volume provides a comprehensive introduction to the theory of electronic motion in molecular processes an increasingly relevant and rapidly expanding segment of molecular quantum dynamics Emphasis is placed on describing and interpreting transitions between electronic states in molecules as they occur typically in cases of reactive scattering between molecules photoexcitation or nonadiabatic coupling between electronic and nuclear degrees of freedom Electron Dynamics in Molecular Interactions aims at a synoptic presentation of some very recent theoretical efforts to solve the electronic problem in quantum molecular dynamics contrasting them with more traditional schemes The presented models are derived from their roots in basic quantum theory their interrelations are discussed and their characteristic applications to concrete chemical systems are outlined This volume also includes an assessment of the present status of electron dynamics and a report on novel developments to meet the current challenges in the field Further this monograph responds to a need for a systematic comparative treatise on nonadiabatic theories of quantum molecular dynamics which are of considerably higher complexity than the more traditional adiabatic approaches and are steadily gaining in importance This volume addresses a broad readership ranging from physics or chemistry graduate students to specialists in the field of theoretical quantum dynamics

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Gillespie,1972-01-01 **The VSEPR Model of Molecular Geometry** Ronald J Gillespie,Istvan Hargittai,2013-03-21
Valence Shell Electron Pair Repulsion VSEPR theory is a simple technique for predicting the geometry of atomic centers in small molecules and molecular ions This authoritative reference was written by Istvan Hartiggai and the developer of VSEPR theory Ronald J Gillespie In addition to its value as a text for courses in molecular geometry and chemistry it constitutes a classic reference for professionals Starting with coverage of the broader aspects of VSEPR this volume narrows its focus to a succinct survey of the methods of structural determination Additional topics include the applications of the VSEPR model and its theoretical basis Helpful data on molecular geometries bond lengths and bond angles appear in tables and other graphics

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