

Molarity Worksheet #1

name _____

1. What does molarity mean?

$$\frac{\text{Number of moles of solute}}{\text{1 liter solution}}$$

2. What is the molarity of a solution that contains 4.53 moles of lithium nitrate in 2.85 liters of solution?

$$\frac{4.53 \text{ mol LiNO}_3}{2.85 \text{ L soln}} = 1.59 \text{ M LiNO}_3$$

3. What is the molarity of a solution that contains 0.00372 moles hydrochloric acid in 2.39×10^{-2} liters of solution?

$$\frac{0.00372 \text{ mol HCL}}{2.39 \times 10^{-2} \text{ L soln}} = 0.156 \text{ M HCL}$$

4. A flask contains 85.5 g $C_{12}H_{22}O_{11}$ (sucrose) in 1.00 liters of solution. What is the molarity?

$$\frac{85.5 \text{ g sucrose} \times 1 \text{ mol sucrose}}{1.00 \text{ L soln} \quad 342.34 \text{ g sucrose}} = 0.250 \text{ M sucrose}$$

5. A beaker contains 214.2 grams osmium (III) fluoride in 0.0673 liters of solution. What is the molarity?

$$\frac{214.2 \text{ g OsF}_3 \times 1 \text{ mol OsF}_3}{0.0673 \text{ L soln} \quad 247.23 \text{ g OsF}_3} = 12.9 \text{ M OsF}_3$$

6. Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution.

$$\frac{1.54 \text{ mol K}_2\text{SO}_4}{0.125 \text{ L soln}} = 12.3 \text{ M K}_2\text{SO}_4$$

Molarity Practice Problems Answer Key With Work

RJ Alexander



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R.R. Bowker's Software for Schools, 1987

General Chemistry James P. Birk, Donald M. Baird, 1994-03

Pocket Guide to Accompany Chemistry & Chemical Reactivity John M. Dekorte, 1999

Resources in Education, 1980

Biology/science Materials Carolina Biological Supply Company, 1991

Elements of General and Biological Chemistry John R. Holum, 1987 This book is a shorter version of the third edition of Fundamentals of General Organic and Biological Chemistry 1986 It incorporates the recommendations of the Task Force on Chemical Education for Health Professions and meets the needs for a basic text in a one term course in chemistry for students aiming for careers in professional health care fields

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Resources in Education, 1980

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