

Half Modern Chemistry Review

CHAPTER 5: THE PERIODIC TABLE

The following pages contain the best (but not all) of the information for this chapter in class.

Focus on this content, but make sure to review other topics, activities, handouts, problems, etc.

If you already take advanced-level AP/IB class, your school is listed for state or P&S. Use that.

**** Use items with the words: examples, another application of this content, etc. ****

COURSES:

- College-level (with general or interdisciplinary) courses: GC 42
- Community-college-level or transfer courses: GC 43
- 5.1: Describe the organization of the periodic table; distinguish between periods and groups. IT 41 42
- 5.2: Use the Periodic Table to identify elements about their properties (chemical and physical), reactivity, energy, and atomic size, of that elements. IT 41 42 43
- 5.3: Identify common compounds, metalloids, and noble gases; identify the representative groups and main ions (groups: 1, 2, 13, 14, 15, 16, 17, 18) IT 41 42
- 5.4: Compare periodic trends of atomic size, ionization energy, and electronegativity. IT 41 42

5.1: PERIODS OF THE PERIODIC TABLE

• Introductory

- periodic law = the law that states that the repeating chemical and physical properties of elements change periodically with the atomic numbers of the elements
- periodic table = an arrangement of the elements in order of their atomic numbers such that elements with similar properties fall in the same column.
- horizontalities = a number of the same number of elements, known as the number, same from left to right in IT (horizontal)
- verticality = set of the same of heavy metals that elements that decrease from top to bottom (atomic number 40) through lanthanides (atomic number 100) on the periodic table

• Chapter Highlights

- This periodic law states that the physical and chemical properties of the elements, and periodic functions, of these elements, change.
- This periodic table is an arrangement of the elements in order of their atomic numbers so that elements with similar properties fall in the same column/group/family.
- The columns in the periodic table are referred to as groups.

5.2: ELECTRON CONFIGURATIONS AND THE PERIODIC TABLE

• Introductory

- shell model = one of the elements of Group 1 of the periodic table (lithium, sodium, potassium, rubidium, cesium, and francium)
- subshell model = one of the elements of Group 2 of the periodic table (beryllium, magnesium, calcium, strontium, barium, and radium)
- representative elements = one of the metals that can use the inner shell electrons using the outer shell to bond.
- metalloid elements = an element in the IT or IB or IVB of the periodic table
- halogens = one of the elements of Group 17 (fluorine, chlorine, bromine, iodine, and astatine); halogens combine with most metals to form salts.

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