

9.1 Assessment

- The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8.
- from the number of electrons lost.
- ite or -ate
- Group 1A metals, 1+; Group 3A (aluminum), 3+; Group 5A nonmetals, 3-
- A monoatomic anion is a single atom with a negative charge; a polyatomic anion is two or more bound atoms with a negative charge.
- K⁺, cation, potassium ion
 - O²⁻ anion, oxide ion
 - Sn²⁺ cation, tin (II)
 - Br⁻ anion, bromide ion
 - Be²⁺ cation beryllium ion
 - Co³⁺ cation, cobalt(III) ion
- NH₄⁺
 - Sn²⁺
 - CrO₄²⁻
 - NO₃

9.2 Assessment

- Write the name of the cation followed by the name of the anion.
- Write the symbol of the cation followed by the symbol of the anion, then use subscripts to balance the charges.
- Write the symbol for the metal ion followed by the formula of the polyatomic ion and balance the charges. Name the cation first followed by the anion.
- BeCl₂
 - Cs₂S
 - NaI
 - SrO
- Cr(NO₂)₃
 - NaClO₄
 - Mg(HCO₃)₂
 - Ca(C₂H₃O₂)₂
- incorrect; charges are not balanced MgSO₄
 -
 - incorrect; charges are not balanced BeCl₂

9.3 Assessment

- Prefixes indicate the number of atoms of each element in a molecule of the compound.

Holt Chemistry Chapter 9 Review Answers

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