

Reaction Rates and Chemical Equilibrium

A. Factors That Affect the Rate of a Reaction

A.1 Effect of Temperature

Test Tube Temperature (3)

Observations (4)

1. 10 °C

Slow fizzing, clear solution after 2 min 38 sec

2. 58 °C

vigorous fizzing/bubbles, clear solution in

Which test tube cleared first (4)?

48 sec

Q1 How did an increase in temperature affect the rate of a reaction?

A.2 Effect of Changing Reactant Concentration

Test Tube	Concentration of HCl	Initial time (2)	Final time (3)	Total time (3) - (2)
1.	1.0 M HCl	0	17 min 5 sec	
2.	2.0 M HCl	0	4 min 38 sec	
3.	3.0 M HCl	0	1 min 20 sec	

Rank the test tubes in order of slowest to fastest. _____ < _____ < _____

Q2 How did an increase in the concentration of HCl cause a difference in the rates of reaction of HCl with Mg?

Lab Report On Reaction Rates And Equilibrium

S Marginson



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