

LABORATORY MANUAL

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**DEPARTMENT OF BIOMEDICAL
ENGINEERING**

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FACULTY OF ENGINEERING & TECHNOLOGY

SRM UNIVERSITY

(UNDER SECTION 3 of UGC ACT 1956)

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Lab Manual Transistor Devices

S Marginson



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web in today s experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is

used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice
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 solution $Ca(OH)_2 \rightleftharpoons Ca^{2+} + 2OH^-$ $5 \times 10^{-2} M$ calcium ions $5 \times 10^{-2} M$ $0 M$ hydroxide ions calculate calcium hydroxide K_{sp} using the
 concentrations of calcium ions and hydroxide ions

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web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the K_{sp} of a slightly
 soluble basic salt and how solubility changes with temperature K_{sp} is another form of equilibrium expression that deals with
 solubility the reaction that will take place is $Ca(OH)_2 \rightleftharpoons Ca^{2+} + 2OH^-$

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 molarity $0.050 M$ volume of filtered $Ca(OH)_2$ solution ml 5.00 5.00 3.71 13.75 equivalence point volume ml oh in saturated $Ca(OH)_2$
 $Ca(OH)_2$ in saturated $Ca(OH)_2$ K_{sp} of $Ca(OH)_2$ average K_{sp} of $Ca(OH)_2$

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web the chemical equation used was $Ca(OH)_2(s) \rightleftharpoons Ca^{2+}(aq) + 2OH^-(aq)$ the solubility product was determined to be $K_{sp} Ca(OH)_2$ from the
 titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium
 hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $Ca_3(PO_4)_2$ is a constant it does not appear explicitly in
 the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore
 $K_{sp} Ca_3(PO_4)_2 = [Ca^{2+}]^3 [PO_4^{3-}]^2$

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web lab report 8 determining of calcium hydroxide the purpose of this was to titrate a saturated $Ca(OH)_2$ solution with a standard
 HCl solution and find the K_{sp} or the OH concentration was found and then the concentration was found and u for this
 experiment a titration of calcium hydroxide with hydrochloric acid

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 calcium hydroxide with hydrochloric acid was investigated determining the K_{sp} of calcium hydroxide $Ca(OH)_2(s) \rightleftharpoons Ca^{2+}(aq) + 2OH^-(aq)$
 for this experiment the titration of calcium

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for Ca(OH)_2 was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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