

Gas Variables

How are the variables that describe a gas related?

Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and amount (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 – Gases in a Nonflexible Container

Experiment A (Adding more gas)



A1

Volume = 2 units

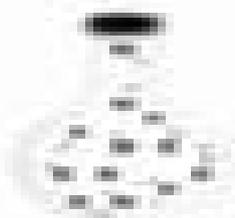
External pressure = 1 atm Internal pressure = 1 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K



A2

Volume = 2 units

External pressure = 1 atm Internal pressure = 2 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K



A3

Volume = 2 units

External pressure = 1 atm Internal pressure = 3 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K

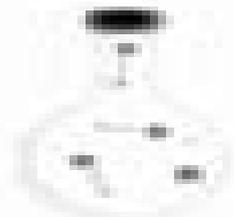
Experiment B (Heating the gas)



B1

Volume = 2 units

External pressure = 1 atm Internal pressure = 1 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 300 K Temperature = 300 K Temperature = 300 K



B2

Volume = 2 units

External pressure = 1 atm Internal pressure = 2 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 400 K Temperature = 400 K Temperature = 400 K



B3

Volume = 2 units

External pressure = 1 atm Internal pressure = 3 atm External pressure = 1 atm
Internal pressure = 1 atm Internal pressure = 1 atm Internal pressure = 1 atm
Temperature = 600 K Temperature = 600 K Temperature = 600 K

Note: Volume is the total number of units rather than three because 3 molecules of gas at these conditions gain sufficient kinetic energy each second ($v = 337$ m/s). These kinetic energies are much larger compared to the space between them than in a liquid or a solid.

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