

Gas Laws – Review Problems

Instructions: Use the Combined Gas Law and Dalton's Law to complete each problem.

Boyle's Law:

$$P_1 V_1 = P_2 V_2$$

Charles' Law:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Gay-Lussac's Law:

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- 1) If the gas inside a volleyball has a pressure of 5.50 atm at a temperature of 309 K, what will the pressure be if the volleyball is reaches a temperature of 279 K?

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \rightarrow \quad \frac{5.50}{309} = \frac{P_2}{279} \quad \rightarrow \quad \boxed{P_2 = 4.97 \text{ atm}}$$

- 2) An inflated balloon contains gas with a volume of 2.5 liters until the pressure on the balloon is changed, increasing the volume of the balloon to 4.7 liters. If the final pressure of the balloon is 2.6 atm, what was the initial pressure?

$$P_1 V_1 = P_2 V_2 \quad P_1 (2.5) = (2.6)(4.7) \quad \boxed{P_1 = 4.89 \text{ atm}}$$

- 3) A car is left outside at night and has a tire pressure of 2.08 atm. In the morning, the temperature is 282 K and the tire pressure has increased to 2.29 atm. What was the initial temperature the previous night?

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \rightarrow \quad \frac{2.08}{T_1} = \frac{2.29}{282} \quad \rightarrow \quad \boxed{T_1 = 256.14 \text{ K}}$$

- 4) A canister is filled with 5.75 liters of gas and heated to a temperature of 600 K. Afterwards, the gas cools to a temperature of 340 K. What was the final volume?

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \rightarrow \quad \frac{5.75}{600} = \frac{V_2}{340} \quad \rightarrow \quad \boxed{V_2 = 3.26 \text{ L}}$$

- 5) A large balloon is filled with 4.65 L of gas at a pressure of 1.21 atm. It is then placed under pressure until the volume decreases to 2.22 L. How much pressure was used?

$$P_1 V_1 = P_2 V_2 \quad (1.21)(4.65) = P_2 (2.22) \quad \boxed{P_2 = 2.53 \text{ atm}}$$

Gas Law Study Guide Answers

Wenbin Ji



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Table of Contents Gas Law Study Guide Answers

1. Understanding the eBook Gas Law Study Guide Answers
 - The Rise of Digital Reading Gas Law Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Law Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Study Guide Answers
 - Personalized Recommendations
 - Gas Law Study Guide Answers User Reviews and Ratings
 - Gas Law Study Guide Answers and Bestseller Lists
5. Accessing Gas Law Study Guide Answers Free and Paid eBooks
 - Gas Law Study Guide Answers Public Domain eBooks
 - Gas Law Study Guide Answers eBook Subscription Services
 - Gas Law Study Guide Answers Budget-Friendly Options
6. Navigating Gas Law Study Guide Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Gas Law Study Guide Answers Compatibility with Devices
 - Gas Law Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Gas Law Study Guide Answers
 - Highlighting and Note-Taking Gas Law Study Guide Answers
 - Interactive Elements Gas Law Study Guide Answers
8. Staying Engaged with Gas Law Study Guide Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Law Study Guide Answers
9. Balancing eBooks and Physical Books Gas Law Study Guide Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Law Study Guide Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Law Study Guide Answers
- Setting Reading Goals Gas Law Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Law Study Guide Answers
- Fact-Checking eBook Content of Gas Law Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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