The ABC's of Gas Laws

A Avogadro's Law

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

- Equal volume of gases, at same pressure and temperature, have an equal number of moles
- Volume is directly proportional to the number of moles
- Explaination of Gay-Lussac's Law of combining volumes.

B Boyle's Law

$$P_1V_1 = P_2V_2$$

- The smaller the volume of the container, the greater the pressure on a gas due to more frequent collisions
- The force of lots of collisions between gas molecules and walls of the container is responsible for the pressure exerted by a gas.
- Constant temperature and number of moles

C Charles's Law

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

- Constant pressure and number of moles
- As the temperature of a gas increases, the gas's volume increases
- As temperature increases, the speed of gas molecules increases, which makes the molecule's collisions with the walls of the container more energetic, thus increasing pressure. If the container is not rigid, then the volume of gas increases (Boyle's Law).

D Dalton's Law of Partial Pressure

$$P_{total} = P_1 + P_2 + \ldots P_n$$
$$P_i = X_i \, P_{total}$$

- **Partial Pressure**: The pressure exerted by each gas in a mixture of ideal gases
- The total pressure of a mixture of ideal gases is the sum of the partial pressures of all the gases in the mixture.

Total pressure of a gas mixture is determined by the number of moles of each gas present in the mixture, called it's **mole fraction**. The sum of the mole fractions for all the gases in a mixture should equal 1.

$$X_i = \frac{n_i}{n_{total}} \qquad X_1 + X_2 = 1$$

E Effusion (Graham's Law)

$$\frac{Rate_A}{Rate_B} = \frac{Time_B}{Time_A} = \sqrt{\frac{Molar\ Mass_B}{Molar\ Mass_A}}$$

- **Effusion** - the relative rates at which a gas in a gas mixture escapes from a container through a hole.
- The larger the molecule, the smaller it's relative rate is. Larger molecules move slower so they cannot travel across a spaces as quickly (**diffusion**) or find the hole to escape as fast (**effusion**).

Diffusion   Effusion

G Gay-Lussac's Law

Hmm.... Where's F?

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- As temperature increases, the pressure increases in a closed container due to more frequent collisions
- Constant volume and number of moles.

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Yi-Tong Ma

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