

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$PV = nRT$

SHOW ALL WORK!!!

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g Na}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = 2.34 \text{ L Cl}_2$

$2.34 \text{ L Cl}_2$



- a. Balance the above reaction.  
 b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$n = x$   
 $V = O_2(x)$

$T_1 = 25 + 273 = 298 \text{ K}$   
 $P_1 = 1.05 \text{ atm}$

$5.53 \text{ g C}_3\text{H}_8 \left| \frac{1 \text{ mol}}{44 \text{ g}} \right| \left| \frac{5 \text{ mol O}_2}{1 \text{ mol C}_3\text{H}_8} \right| = 0.628 \text{ moles}$   
 $V = 14.6 \text{ L}$

$14.6 \text{ L}$

3. Potassium permanganate,  $\text{KMnO}_4$ , is produced commercially by oxidizing aqueous potassium manganate,  $\text{K}_2\text{MnO}_4$ .



- a. Balance the above reaction.  
 b. What volume of  $\text{Cl}_2(g)$ , measured at STP, is needed to produce 10.0 g of  $\text{KMnO}_4$ ?

$V = \frac{nRT}{P}$

$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mol}}{158 \text{ g}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol Cl}_2} \right| = 0.709 \text{ L Cl}_2$

$0.709 \text{ L}$

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.  
 b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mol}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{1 \text{ mol}} \right| = 0.204 \text{ mol}$   
 Lim Reagent

$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mol}}{18 \text{ g}} \right| \left| \frac{2 \text{ mol NH}_3}{3 \text{ mol}} \right| = 0.381 \text{ mol}$

$24 + 273 = 297 \text{ K}$   
 $752 / 760 = 0.989 \text{ atm}$

$5.03 \text{ L}$

$V = \frac{nRT}{P}$

$V = \frac{(0.204 \text{ mol} + 0.082 \text{ mol}) (0.082 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}) (297 \text{ K})}{0.989 \text{ atm}}$

# Gas Stoichiometry Answer Key

**Albert A Gayle**



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## **Table of Contents Gas Stoichiometry Answer Key**

1. Understanding the eBook Gas Stoichiometry Answer Key
  - The Rise of Digital Reading Gas Stoichiometry Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Stoichiometry Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Stoichiometry Answer Key
  - Personalized Recommendations
  - Gas Stoichiometry Answer Key User Reviews and Ratings
  - Gas Stoichiometry Answer Key and Bestseller Lists
5. Accessing Gas Stoichiometry Answer Key Free and Paid eBooks
  - Gas Stoichiometry Answer Key Public Domain eBooks
  - Gas Stoichiometry Answer Key eBook Subscription Services

- Gas Stoichiometry Answer Key Budget-Friendly Options
- 6. Navigating Gas Stoichiometry Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Stoichiometry Answer Key Compatibility with Devices
  - Gas Stoichiometry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Stoichiometry Answer Key
  - Highlighting and Note-Taking Gas Stoichiometry Answer Key
  - Interactive Elements Gas Stoichiometry Answer Key
- 8. Staying Engaged with Gas Stoichiometry Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Stoichiometry Answer Key
- 9. Balancing eBooks and Physical Books Gas Stoichiometry Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Stoichiometry Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Stoichiometry Answer Key
  - Setting Reading Goals Gas Stoichiometry Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Stoichiometry Answer Key
  - Fact-Checking eBook Content of Gas Stoichiometry Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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