

Gas Stoichiometry Worksheet

Name:

Solve all the following gas law problems. Show all work, answers are given at the end of the problem.

Molar Volume

1. Calculate the number of moles contained in 550.mL of carbon dioxide at STP. (0.0246mol)

2. Calculate the mass of 1.50 L of CH_4 at STP. (1.07g)

3. Calculate the volume in liters of 50.0 grams of nitrogen dioxide at STP. (24.3L)

Gas Stoichiometry

4. Given the following reaction: $\text{NH}_4\text{NO}_2 (\text{s}) \rightarrow \text{N}_2 (\text{g}) + 2\text{H}_2\text{O} (\text{g})$ How many liters of nitrogen gas is produced if 50.0L of water is produced at STP? (25.0L)

5. Given the following reaction: $\text{CuO} (\text{s}) + \text{H}_2 (\text{g}) \rightarrow \text{Cu} (\text{s}) + \text{H}_2\text{O} (\text{g})$ If 250. L of hydrogen gas are used to reduce copper (II) oxide at STP, what mass of copper is to be expected? (709g)

6. How many grams of solid lithium must be added to liquid water in order to obtain 15.0 L of hydrogen gas at STP? (9.29g) $2\text{Li} + 2\text{H}_2\text{O} \rightarrow 2\text{LiOH} + \text{H}_2$

Gas Stoichiometry Practice Test

Richard Bailey



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Table of Contents Gas Stoichiometry Practice Test

1. Understanding the eBook Gas Stoichiometry Practice Test
 - The Rise of Digital Reading Gas Stoichiometry Practice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Stoichiometry Practice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Stoichiometry Practice Test
 - Personalized Recommendations
 - Gas Stoichiometry Practice Test User Reviews and Ratings
 - Gas Stoichiometry Practice Test and Bestseller Lists
5. Accessing Gas Stoichiometry Practice Test Free and Paid eBooks

- Gas Stoichiometry Practice Test Public Domain eBooks
 - Gas Stoichiometry Practice Test eBook Subscription Services
 - Gas Stoichiometry Practice Test Budget-Friendly Options
6. Navigating Gas Stoichiometry Practice Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Stoichiometry Practice Test Compatibility with Devices
 - Gas Stoichiometry Practice Test Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Stoichiometry Practice Test
 - Highlighting and Note-Taking Gas Stoichiometry Practice Test
 - Interactive Elements Gas Stoichiometry Practice Test
 8. Staying Engaged with Gas Stoichiometry Practice Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Stoichiometry Practice Test
 9. Balancing eBooks and Physical Books Gas Stoichiometry Practice Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Stoichiometry Practice Test
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Gas Stoichiometry Practice Test
 - Setting Reading Goals Gas Stoichiometry Practice Test
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Gas Stoichiometry Practice Test
 - Fact-Checking eBook Content of Gas Stoichiometry Practice Test
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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