

Gas Stoichiometry problems

1. Calculate the mass of hydrogen peroxide needed to obtain 0.460 L of oxygen gas at STP.



$$0.460 \text{ L O}_2 \left| \frac{1 \text{ mole O}_2}{22.4 \text{ L O}_2} \right| \left| \frac{2 \text{ moles H}_2\text{O}_2}{1 \text{ mole O}_2} \right| \left| \frac{34.0 \text{ g H}_2\text{O}_2}{1 \text{ mole H}_2\text{O}_2} \right| = \boxed{1.40 \text{ g}}$$

Given:

$V = 0.460 \text{ L}$
 $T = 273 \text{ K}$
 $P = 1 \text{ atm}$
 $n = ?$

2. Sodium azide (NaN_3) is used in some air bags. The impact of a collision causes the sodium azide to decompose and form solid Sodium and nitrogen gas. Calculate the volume of nitrogen gas generated at 80.0°C and 823 mm Hg by the decomposition of 60.0 g of NaN_3 .

Balanced equation:



Given:

$T = 353 \text{ K}$
 $P = 823 \text{ mm Hg}$
 $V_{\text{N}_2} = ?$

$$60.0 \text{ g NaN}_3 \left| \frac{1 \text{ mole NaN}_3}{65.0 \text{ g NaN}_3} \right| \left| \frac{3 \text{ moles N}_2}{2 \text{ moles NaN}_3} \right| = 1.38 \text{ moles N}_2$$

$$V_{\text{N}_2} = \frac{nRT}{P} = \frac{(1.38 \text{ moles}) (0.0821 \text{ L atm / mol K}) (353 \text{ K})}{1.08 \text{ atm}}$$

$$V = 37.0 \text{ L}$$

$$823 \text{ mm Hg} \left| \frac{1 \text{ atm}}{760 \text{ mm Hg}} \right| = 1.08 \text{ atm}$$

3. When heated, solid potassium chlorate (KClO_3) decomposes to form solid potassium chloride and oxygen gas. If 20.8 g of potassium chlorate decomposes, how many liters of oxygen gas will form at STP?

Balanced equation:



Given:

$m = 20.8 \text{ g}$
 $T = 273 \text{ K}$
 $P = 1 \text{ atm}$
 $V = ?$
 $\boxed{\text{STP}}$

$$20.8 \text{ g KClO}_3 \left| \frac{1 \text{ mole KClO}_3}{122.55 \text{ g KClO}_3} \right| \left| \frac{3 \text{ moles O}_2}{2 \text{ moles KClO}_3} \right| = 0.801 \text{ moles O}_2$$

$$V_{\text{O}_2} = \frac{nRT}{P} = \frac{(0.801 \text{ moles}) (0.0821 \text{ L atm / mol K}) (273 \text{ K})}{1 \text{ atm}}$$

$$\boxed{V_{\text{O}_2} = 5.72 \text{ Liters}}$$

OR - Since at STP - do in one step.

$$20.8 \text{ g KClO}_3 \left| \frac{1 \text{ mole KClO}_3}{122.55 \text{ g KClO}_3} \right| \left| \frac{3 \text{ moles O}_2}{2 \text{ moles KClO}_3} \right| \left| \frac{22.4 \text{ L O}_2}{1 \text{ mole O}_2} \right| = \boxed{5.70 \text{ L}}$$

Gas Stoichiometry Worksheet Answers With Work

Wolfgang Guggemos



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