

Procedure for making NaOH-Acid (for every lab group of 3-4 people):

- Using your calculations from question #4 above, measure out the correct amount of solid NaOH powder (grams) into the assigned cup to make a 0.1 M solution. **Be careful not to include the mass of the cup in your measurement!**
- Transfer the powder into the assigned "mixing cup".
- Add water into the "mixing cup" until you have 100 mL of solution. Stir with a spoon or stick.
- Observe and taste the solutions you have made. You can have one "designated taster" or you can pour a little into separate little cups for each group member to taste. Record in data table.
- Repeat steps 1-3 in order to make the 0.4 M and 0.7 M solutions.
- Clean-Up. Throw away cups that you drink out of, but save the mixing & measuring cups.

Data / Observations:

Observations (color, taste, etc.)	0.1 M solution	0.4 M solution	0.7 M solution
	Answers will vary	Answers will vary	Answers will vary

Calculations / Analysis:

- Which concentration that you tasted was closest to the ideal concentration of NaOH-Acid? What was wrong with each of the other solutions that you made?

Answers will vary. Some solutions will not have enough NaOH-Acid for your taste, while others will have too much.

- What was the solute used in this lab? What was the solvent?

Solute = NaOH-Acid powder Solvent = water

- What is molarity?

Molarity is a measure of the number of moles of solute per liter of solution.

Questions for Discussion: Do your best show the formulae of your work, and **highlight** your answer!

- Write the formulae for the 4 different ways to calculate concentrations:

- Molarity

$$M = \frac{\text{moles}}{\text{volume}}$$

- grams/Liter

$$\frac{\text{grams of solute}}{\text{volume of solution}}$$

- % composition

$$\frac{\text{grams of solute}}{\text{grams of solution}} \times 100$$

- Parts per million

$$\frac{\text{grams of solute}}{\text{grams of solution}} \times 10^6$$

Determine the following concentrations, using the 4 different methods:

- If 1.00 moles of NaOH was dissolved in enough water to make 1.00 L of solution:

- Molarity?

$$M = \frac{1.00 \text{ mol}}{1.00 \text{ L}}$$

$$M = 1.000 \text{ M NaOH}$$

- grams/Liter?

$$\frac{40.0 \text{ g NaOH}}{1.00 \text{ L}}$$

$$= 40.0 \text{ g/L}$$

- % composition?

$$\frac{40.0 \text{ g NaOH}}{40.0 \text{ g}} \times 100$$

$$= 100\% \text{ NaOH}$$

- Parts per million?

$$\frac{40.0 \text{ g NaOH}}{40.0 \text{ g}} \times 10^6$$

$$= 1,000,000 \text{ ppm}$$

- If 20 grams of NaOH was dissolved in enough water to make a 250 mL NaOH solution:

- Molarity?

$$M = \frac{0.50 \text{ mol}}{0.250 \text{ L}}$$

$$M = 2.0 \text{ M NaOH}$$

- grams/Liter?

$$\frac{20 \text{ g NaOH}}{0.250 \text{ L}}$$

$$= 80 \text{ g/L}$$

- % composition?

$$\frac{20 \text{ g NaOH}}{250 \text{ g}} \times 100$$

$$= 8\% \text{ NaOH}$$

- Parts per million?

$$\frac{20 \text{ g NaOH}}{250 \text{ g}} \times 10^6$$

$$= 80,000 \text{ ppm}$$

Kool Aid Lab Answer Key

Bright Thinker



Kool Aid Lab Answer Key:

Laboratory Exercises for Sensory Evaluation Harry T. Lawless, 2012-12-12 Laboratory exercises are a necessary part of science education. They enable students to better understand the principles discussed in lectures and provide them with hands-on experience of the practical aspects of scientific research. The purpose of this book is to provide students and instructors with a time-tested set of lab exercises that illustrate the common sensory tests and/or sensory principles used in evaluation of foods, beverages, and consumer products. The appendices will also include a set of simple problem sets that can be used to teach and reinforce basic statistical tests. Approximately twenty years ago, the Sensory Evaluation Division of the Institute of Food Technologists sponsored the preparation of a set of exercises titled *Guidelines for Laboratory Exercises for a Course in Sensory Evaluation of Foods*, edited by one of the co-authors, Heymann. This book will provide additional materials from the second author, Lawless, as well as other instructors in a uniform format that can be easily adopted for course use. Most importantly, the lab exercises will complement the flagship textbook in the field, *Sensory Evaluation of Foods: Principles and Practices 2E*, also by Lawless and Heymann, and published by Springer. Possible course adoption of the main text along with the lab manual should enhance the sales of these materials.

Tried and True National Science Teachers Association, 2010 A compilation of popular *Tried and True* columns originally published in *Science Scope*, this new book is filled with teachers' best classroom activities, time-tested, tweaked, and engaging. These ageless activities will fit easily into your middle school curriculum and serve as go-to resources when you need a tried-and-true lesson for tomorrow. From publisher description.

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