

Study Guide Term 3 Chemistry

Chemical Bonding

Study guide for grade 9 Unit 2 - Chemical bonding - Monday June 10th.

key concepts and prepare for assessments - Worksheets and other resources are attached on MB.

- Definition of chemical bonding
- Importance of chemical bonding in understanding the properties of matter
- Formation of ions
- Definition, properties, examples of ionic bonding
- Definition, properties, examples of covalent bonding
- Definition, properties, examples of metallic bonding
- Losing, gaining, and sharing of electrons.
- Naming ionic compounds with mono and polyatomic ions, and transition metals, naming molecular compounds.
- Electron dot diagram, rules for bonding, octet, and duet rules.
- Introduction to Lewis structures, rules for drawing Lewis structures. Examples of Lewis structures for simple molecules and ions

Practice problems: Extra worksheet.

- Manage bac under "End of year exam Chemistry".

Review questions.

- A list of review questions covering key concepts from each section to help students gauge their understanding in files on Manage Bac under task "End of year exam Chemistry".

Chemical bonding is the process by which atoms form chemical compounds by sharing or transferring electrons. Understanding chemical bonding is crucial in chemistry, as it explains the stability and reactivity of different substances.

Why do atoms lose, gain or share electrons (why do atoms bond)? to become more stable.

Duet Rule

The duet rule states that atoms of hydrogen (H) and helium (He) are most stable when they have 2 electrons in their valence shell.

Octet

Rule

The octet rule states that atoms, other than hydrogen and helium, are most stable when they have 8 electrons in their valence shell.

Determine whether the following atoms have a stable duet or octet configuration:

- Lithium (Li)
- Neon (Ne)
- Fluorine (F)

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