

Honors Chemistry  
Equilibrium Worksheet

**Equilibrium: ICE box practice problems**

1. In a 10.0L vessel at 1000K, 0.250 mol  $\text{SO}_2(\text{g})$  and 0.200 mol  $\text{O}_2(\text{g})$  react to form 0.162 mol  $\text{SO}_3(\text{g})$  at equilibrium. What is the  $K_c$  at 1000K for this reaction?

Balanced equation:  $\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \leftrightarrow \text{SO}_3(\text{g})$

	$\text{SO}_2$	+	$\text{O}_2$	$\leftrightarrow$	$\text{SO}_3$
I					
C					
E					

2. An aqueous solution of 0.00325M  $\text{Fe}^{3+}$  was reacted with 0.0020M  $\text{SCN}^-$  at 25°C. Calculate the value of the equilibrium constant for this reaction when the equilibrium concentration of  $\text{FeSCN}^{2+}$  was determined to be 0.000667M

	$\text{Fe}^{3+}$	+	$\text{SCN}^-$	$\leftrightarrow$	$\text{FeSCN}^{2+}$
I					
C					
E					

3. At a certain temperature, 2.0 mole  $\text{NH}_3$  is introduced into a 2.0 L container, and the  $\text{NH}_3$  partially dissociates into by the following reaction:



At equilibrium, 1.0 mol  $\text{NH}_3$  remains, what is the value of  $K$  for this reaction?

	$2\text{NH}_3(\text{g})$	$\leftrightarrow$	$\text{N}_2(\text{g})$	+	$3\text{H}_2(\text{g})$
I					
C					
E					

# Equilibrium Ice Box Practice Problems

**SJ Ball**



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