

Dynamic Equilibrium

At what point is a reversible reaction “completed?”

Why?

Most of the reactions we have studied year have been “forward” reactions—once the reactant has been changed into the product it stays that way. The reaction “stops” when the limiting reactant runs out.

However, in a reversible reaction (one that takes place in both directions), the product can turn back into reactant in a continuous cycle that never stops. So...

When can we say the reaction is “finished?”

Can we predict how much reactant and how much product will be present when the reaction is “done?”

Model 1 – A Reversible Reaction



1. What is the significance of the double arrow in Model 1?
2. Imagine the reaction in Model 1 starts with 100 molecules of A and zero molecules of B in a 1 liter container. Explain why the concentration of reactant A will never reach zero.
3. Imagine the reaction in Model 1 starts with zero molecules of A and 100 molecules of B in a 1 liter container. Will the concentration of product B increase, decrease, or remain the same as the reaction proceeds?

Explain.

Equilibrium Pogil Answers

Terry C. Jones



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