

Equilibrium Constant - Worksheet 1_ANSWERS

1. Write the equilibrium expression for the following reactions:



$$K = \frac{[U]^u [V]^v}{[W]^w [X]^x}$$



$$K = \frac{[H_2][CO_2]}{[H_2O][CO]}$$



$$K = \frac{[Cl_2][CO]}{[COCl_2]}$$



$$K = \frac{[HCl]^2}{[H_2][Cl_2]}$$



$$K = \frac{[NO][CO_2]}{[NO_2][CO]}$$



$$K = \frac{[Zn^{2+}]}{[Ag^+]^2}$$

*Note: solids are not included



$$K = \frac{[H_2][C_2H_4]}{[C_2H_6]}$$

2. Consider the following equilibrium reaction:



The K_{eq} for this reaction at 25°C is 1.02. At equilibrium the concentration of HBr is 0.50 mol/L. Assuming H_2 and Br_2 are present in equal amounts, calculate the concentration of H_2 at equilibrium.

$$[H_2] = 0.50 \text{ mol/L}$$

Equilibrium Questions With Answers

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Table of Contents Equilibrium Questions With Answers

1. Understanding the eBook Equilibrium Questions With Answers
 - The Rise of Digital Reading Equilibrium Questions With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Equilibrium Questions With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Equilibrium Questions With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Equilibrium Questions With Answers
 - Personalized Recommendations
 - Equilibrium Questions With Answers User Reviews and Ratings
 - Equilibrium Questions With Answers and Bestseller Lists
5. Accessing Equilibrium Questions With Answers Free and Paid eBooks

- Equilibrium Questions With Answers Public Domain eBooks
 - Equilibrium Questions With Answers eBook Subscription Services
 - Equilibrium Questions With Answers Budget-Friendly Options
6. Navigating Equilibrium Questions With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Equilibrium Questions With Answers Compatibility with Devices
 - Equilibrium Questions With Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Equilibrium Questions With Answers
 - Highlighting and Note-Taking Equilibrium Questions With Answers
 - Interactive Elements Equilibrium Questions With Answers
 8. Staying Engaged with Equilibrium Questions With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Equilibrium Questions With Answers
 9. Balancing eBooks and Physical Books Equilibrium Questions With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Equilibrium Questions With Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Equilibrium Questions With Answers
 - Setting Reading Goals Equilibrium Questions With Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Equilibrium Questions With Answers
 - Fact-Checking eBook Content of Equilibrium Questions With Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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