

Name: _____

Period: _____



Mixed Gas Law Calculations (Boyle's, Charles', and Gay-Lussac's)

Instructions: Solve each of the following problems **and** identify the law.

- 1) 2.00 L of a gas is at 740.0 mmHg pressure. What is its volume at standard pressure?

Law: Boyles

(Ans: 1.95 L)

std pressure:
1 atm = 101.325 mmHg

$$P_1 \cdot V_1 = P_2 \cdot V_2$$

$$740.0 \cdot 2.00 = 101.325 \cdot V$$

$$V = 1.95$$

- 2) A gas is collected and found to fill 2.85 L at 25.0°C. What will be its volume at standard temp? Law: Charles

(Ans: 2.61 L)

273.15 K
standard temp

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{2.85 \text{ L}}{298.15 \text{ K}} = \frac{V}{273.15 \text{ K}} = 2.61$$

- 3) A chemist collected 56.1 mL of gas in an open manometer. The next day, the chemist noted that the volume had changed to 57.9 mL and the barometer reading was 99.4 kPa. The temperature had not changed. What had been the barometer reading on the previous day when the gas was collected?

Law: Boyles

(Ans: 1.01 atm)

$$P_1 \cdot V = P_2 \cdot V$$

$$56.1 \text{ mL} \cdot P = 99.4 \text{ kPa} \cdot 57.9 \text{ mL} = 101$$

- 4) A glass sphere is filled to full volume with a gas. The pressure of the gas inside the sphere is 30.0 atm, and the temperature is 25.0°C. The sphere is taken outside on a cold day. The temperature of the gas decreases to 10.0°C. What is the new pressure of the gas? Assume that the volume is constant.

Law: G-L

(Ans: 28.5 atm)

$$\frac{P}{T} = \frac{P}{T}$$

$$\frac{30.0 \text{ atm}}{298^\circ\text{C}} = \frac{P}{10^\circ\text{C}} = 28.5$$

- 5) The gas pressure in an aerosol canister is 151.6 kPa at 25.0°C, what would the pressure be inside the canister at 300.0°C?

Law: G-L

(Ans: 2.88 atm)

$$\frac{P}{T} = \frac{P}{T}$$

$$\frac{151.6 \text{ kPa}}{298} = \frac{P}{300} = 2.88$$

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LM Reder



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