

MIXED Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.



STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: W5.1 work)

Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.

$$\% \text{ Al} = \frac{\text{total mass Al}}{\text{molar mass Al}_2(\text{SO}_4)_3} \times 100 \% = \frac{53.96 \text{ g}}{342.17 \text{ g}} \times 100 \% = 15.77 \%$$

STEP 4: Complete the following problems using **a factor chart with units & chemicals**. These problems could require ONE ratio, TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

PROBLEMS

1) How many **grams of calcium sulfate** are produced from reacting **2.59 moles of aluminum sulfate**?

$$2.59 \text{ mol Al}_2(\text{SO}_4)_3 \times \frac{3 \text{ mol CaSO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{136.15 \text{ g CaSO}_4}{1 \text{ mol CaSO}_4} = 1057.89 \text{ g CaSO}_4$$

moles A \longrightarrow moles B \longrightarrow grams B

Mixed Stoichiometry Answers

DP Hallahan



Mixed Stoichiometry Answers:

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Table of Contents Mixed Stoichiometry Answers

1. Understanding the eBook Mixed Stoichiometry Answers
 - The Rise of Digital Reading Mixed Stoichiometry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Stoichiometry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mixed Stoichiometry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Stoichiometry Answers
 - Personalized Recommendations
 - Mixed Stoichiometry Answers User Reviews and Ratings
 - Mixed Stoichiometry Answers and Bestseller Lists
5. Accessing Mixed Stoichiometry Answers Free and Paid eBooks
 - Mixed Stoichiometry Answers Public Domain eBooks
 - Mixed Stoichiometry Answers eBook Subscription Services
 - Mixed Stoichiometry Answers Budget-Friendly Options
6. Navigating Mixed Stoichiometry Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Mixed Stoichiometry Answers Compatibility with Devices
 - Mixed Stoichiometry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Stoichiometry Answers
 - Highlighting and Note-Taking Mixed Stoichiometry Answers
 - Interactive Elements Mixed Stoichiometry Answers
 8. Staying Engaged with Mixed Stoichiometry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Stoichiometry Answers
 9. Balancing eBooks and Physical Books Mixed Stoichiometry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Stoichiometry Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mixed Stoichiometry Answers
 - Setting Reading Goals Mixed Stoichiometry Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mixed Stoichiometry Answers
 - Fact-Checking eBook Content of Mixed Stoichiometry Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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