

## Section Quiz: Introduction to Stoichiometry

In the space provided, write the letter of the term or phrase that best completes each sentence or best answers each question.

- C 1. Stoichiometry is the branch of chemistry that deals with elements in compounds and with reactants and products in chemical reactions, focusing on
- bonding.
  - energy transfers.
  - mass relationships.
  - physical characteristics.
- A 2. Reaction stoichiometry is based on chemical equations and
- the law of conservation of mass.
  - the ideal gas law.
  - the law of conservation of momentum.
  - Avogadro's number.
- C 3. The number of significant figures in an answer to a stoichiometry problem is determined only by
- the number of significant figures in the mole ratio used to solve the problem.
  - the number of decimal places in the molar masses of substances in the chemical equation.
  - the number of significant figures of any measured quantities in the problem.
  - the number of significant figures in the molar masses of substances in the chemical equation.
- D 4. The chemical equation  $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$  can be interpreted correctly as
- 1 molecule of  $\text{P}_4\text{O}_{10} + 6$  molecules of  $\text{H}_2\text{O}$  form 4 molecules of  $\text{H}_3\text{PO}_4$ .
  - 1 mol of  $\text{P}_4\text{O}_{10} + 6$  mol of  $\text{H}_2\text{O}$  reacts to form 4 mol of  $\text{H}_3\text{PO}_4$ .
  - 1 g of  $\text{P}_4\text{O}_{10} + 6$  g of  $\text{H}_2\text{O}$  produces 4 g of  $\text{H}_3\text{PO}_4$ .
  - Both (a) and (b)
- C 5. How many mole ratios can be obtained from the following chemical equation?

- 1
- 4
- 12
- 16



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**B Lingard**



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