



Messenger Chapter 20

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38 grams of gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

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use Avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at STP. $n = \frac{V}{V_m} = \frac{0.0896 \text{ l}}{22.4 \text{ l/mol}} = 0.0040 \text{ moles}$. Experimental molar mass of butane is 58 g/mol. The molecular formula of butane is C_4H_{10} .

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in this experiment the moles and mass will be required to determine the molar mass of butane in a butane lighter. experimental to begin the experiment the mass of a butane lighter was measured before using the butane gas to deplete the water to the 80ml mark on a graduated cylinder

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web although the most convenient gas for this is butane other gases may be available in similar small cylinders the collection of a gas sample and the weighing of the gas cylinder before and after this should take about 5 10 minutes as a demonstration

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