

Determine the mass number of the other isotope in the sample of antimony.

$$A_r = \frac{(\text{mass isotope 1} \times \%) + (\text{mass isotope 2} \times \%)}{100}$$

$$121.8 = \frac{(121 \times 60) + (? \times 40)}{100}$$

$$\frac{(121.8 \times 100) - (121 \times 60)}{40} = ?$$

mass number of the other antimony isotope = 123 (1)

(d) Antimony chloride, $SbCl_3$, exists as simple covalent molecules.

A 'dot-and-cross' diagram of $SbCl_3$ is shown below.

Same as $AlCl_3$
(in notes!!)



4 pairs of e⁻s

6 p_o? (p_o?)

equivalent?

(i) Predict the shape of a molecule of $SbCl_3$.

Explain your answer.

name of shape Pyramidal

4 pairs of electrons in valence shell

3 bonds, 1 lone pair, 7 valence e⁻ on Sb

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