

Name: _____ Block: _____ Date: _____

PACKET #1

Unit 1: Chemistry of Life, Part I

*Reading: BSCS Text Chapter 1, 2.9, 2.10**Learning Objectives:**Topic 1: Basic Chemistry and Bonding*

1. Explain the relationships among atoms and molecules (1.1).
2. Describe and diagram the structure of an atom, including protons, neutrons, electrons (1.2).
3. Use atomic number on the periodic table to determine number of protons and electrons in an atom (class).
4. Explain why atoms form covalent bonds and use a Bohr model/HONC rule to predict bonding behavior(1.4).
5. Define and draw an ion, and provide examples (1.4).
6. Explain why atoms form ionic bonds. Given the number of valence electrons in an atom, predict whether the atom is likely to participate in ionic bonding.

Topic 2: Properties of water

7. Explain why water molecules are considered polar and how this helps hydrogen "bonds" form (1.4)
8. How is an **inter**-molecular force (like a hydrogen "bond") different from a **intra**molecular force (like a covalent bond)?
9. *Explain* the following properties of water (including how these are biologically relevant): high surface tension, high specific heat, solid water is less dense than liquid water, water as the "universal solvent," and capillary action. Along the way, be sure to address cohesion and adhesion (p. 189, teacher notes).
10. Differentiate between acids and bases on a pH scale.

Key Terms/Concepts

Molecules	Covalent Bond	Intermolecular force
Atoms	Ionic Bond	Intramolecular force
Element	Ion	Cohesion
Subatomic particles	Ionization	Adhesion
Proton	Hydrogen bond	Capillary action
Neutron	Polar molecule	Surface tension
Electron	Atomic number	Specific heat
Solution	Solute	Salt
Acid	Base	Solvent

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