



| Events                              |   | Problem                    | Offenders       |
|-------------------------------------|---|----------------------------|-----------------|
| <input type="checkbox"/>            | 🔊 | Excess Networked Flows     | 148000001018048 |
| <input type="checkbox"/>            | 🔊 | Excess Networked Flows     | 148000001018048 |
| <input type="checkbox"/>            | 🔊 | TCP Reset Flood            | 3000001         |
| <input type="checkbox"/>            | 🔊 | Invalid Src/Dst Flows      | 3000001         |
| <input type="checkbox"/>            | 🔊 | Invalid Src Flows          | 3000001         |
| <input checked="" type="checkbox"/> | 🔊 | Invalid Dst Flows          | 3000001         |
| <input type="checkbox"/>            | 🔊 | Invalid Src/Dst Flows      | 3000001         |
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| <input type="checkbox"/>            | 🔊 | Invalid Src Flows          | 148000001120148 |
| <input type="checkbox"/>            | 🔊 | TCP Reset Flood (Reversed) | 17124150100     |
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### Invalid Dst Flows via Router

Today 11:30 AM IST

Event Summary

|                |                                   |                             |                            |
|----------------|-----------------------------------|-----------------------------|----------------------------|
| Severity       | 🟡                                 | Status                      | Open                       |
| Volume         | 6000000                           | Packets                     | 000                        |
| Hits           | 100                               | Problem Class               | Invalid Flows              |
| Algorithm Used | Router Aggregation                | Event Aggregation Threshold | ICM1_FLOW1 greater than... |
| Time           | 2024-09-01 10:00-2024-09-01 11:30 | Event ID                    | 00114                      |

100%

Source Occupancy

100%

Destination Occupancy

Src IP / Src Port

Dst IP / Dst Port

# Manageengine Netflow Analyzer Manual

**Tao Wei**



## **Manageengine Netflow Analyzer Manual:**

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web 1 the aufbau principle states that electrons are always placed in the lowest energy sublevel that is available 2 the pauli exclusion principle states that each orbital may contain a maximum of 2 electrons electrons must have opposite spins 3

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web an atomic orbital can hold a maximum of 2 electrons each with opposite spins an atomic orbital can hold a maximum of 6 electrons each with the same spin an atomic orbital can hold a minimum of 2 electrons each with opposite spins multiple choice edit please save your changes before editing any questions

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web orbital diagrams are pictorial representations of the electron configuration showing the individual orbitals and the pairing arrangement of electrons we start with a single hydrogen atom atomic number 1 which consists of one proton and one electron

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web the four different types of orbitals s p d and f have different shapes and one orbital can hold a maximum of two electrons the p d and f orbitals have different sublevels thus can hold more electrons as stated the electron configuration of each element is unique to its position on the periodic table

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web orbitals and electron configurations electron configurations homework 1 write the full electron configuration and orbital diagram for the elements c mo and sb 2 write the abbreviated electron configuration for c mo and sb 3 how many valence electrons are there for c and sb omit transition metals for this type of question 4

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web the electronic configuration of an atom in the quantum mechanical model is stated by listing the occupied orbitals in order of filling with the number of electrons in each orbital indicated by superscript in this notation the electronic configuration of sodium would be  $1s^2 2s^2 2p^6 3s^1$  distributed in the orbitals as

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web the orbital diagram of helium is  $1s^2$  the spin of an electron is specified by a fourth quantum number called the the two

possible values of  $m_s$  are spin up and spin down the three principles that are needed to specify orbital occupancy for the electrons are Pauli exclusion principle no orbital may have more than two electrons and they must have Pauli exclusion principle no orbital may have more than two electrons

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web the electron configuration and the orbital diagram are following hydrogen is the noble gas helium which has an atomic number of 2 the helium atom contains two protons and two electrons the first electron has the same four quantum numbers as the hydrogen atom electron  $n=1, l=0, m_l=0, m_s=1/2$

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web chemistry questions and answers a write the electron configuration draw the orbital diagram determine the distinguishing electron and determine the 4 quantum numbers for the distinguishing electron of the element magnesium mg write electron configurations as  $1s^2 2s^2 2p^6 3s^2 3p^6$  with a space between each entry and no superscripts or  $n$

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web the easiest way to create electron configurations is using an electron configuration table which is a way of writing down the various orbitals available to electrons this table is easy to remember and it makes it possible to generate the electron configuration table for

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web this problem has been solved you will get a detailed solution from a subject matter expert that helps you learn core concepts question write the complete electron configuration and orbital diagram for each of the following enter up to indicate an upwards pointing arrow down to indicate a downwards pointing arrow up down to indicate two

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web determine element from electron configuration determine number of electrons in an element s orbital s element groups to use select at least one s block alkali metals and alkaline earth metals p block nonmetals metalloids and other metals d

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web chemistry chemistry questions and answers part a write the full electron configuration for the following  $Si^3+$  part b write the noble gas abbreviated electron configuration for the following  $Nb^4+$  part c draw the orbital diagram for the following excited  $Ca^2+$  determine the number of unpaired electrons from your previous answer