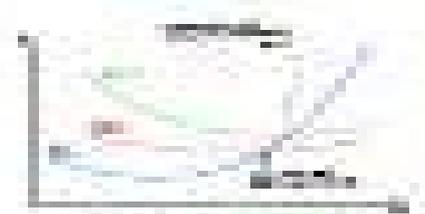


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Introduction: This course will cover the following topics:
 - **Thermodynamics:** Laws of thermodynamics, internal energy, enthalpy, entropy, Gibbs free energy.
 - **Chemical Equilibrium:** Equilibrium constants, Le Chatelier's principle, reaction quotients.
 - **Electrochemistry:** Redox reactions, galvanic cells, standard electrode potentials, Nernst equation.
 - **Chemical Kinetics:** Reaction rates, rate laws, activation energy, Arrhenius equation, reaction mechanisms.

Thermodynamics



First Law of Thermodynamics: Energy is conserved. $\Delta U = q + w$
Second Law of Thermodynamics: The entropy of the universe increases for a spontaneous process. $\Delta S_{univ} > 0$
Gibbs Free Energy: $\Delta G = \Delta H - T\Delta S$. $\Delta G < 0$ indicates a spontaneous process.



Equilibrium Constants: $K = \frac{[products]}{[reactants]}$
Le Chatelier's Principle: A system at equilibrium will shift to counteract any change in conditions.

Electrochemistry: Redox reactions, galvanic cells, standard electrode potentials, Nernst equation.



Chemical Kinetics: Reaction rates, rate laws, activation energy, Arrhenius equation, reaction mechanisms.



Reaction Rate: $\text{Rate} = -\frac{1}{2} \frac{d[H_2O_2]}{dt} = \frac{d[H_2O]}{dt} = \frac{d[O_2]}{dt}$

Rate Law: $\text{Rate} = k[H_2O_2]^2$



Arrhenius Equation: $\ln k = \ln A - \frac{E_a}{RT}$



Reaction Mechanism: A series of elementary steps that describe the overall reaction.

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Kathleen Armour



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