

GENERAL CHEMISTRY 2
CHAPTER TEST 2

Name: _____
Section: _____

Score: _____
Date: _____

Direction: Choose the letter of the correct answer. (*Strictly no erasures on your answer sheets*).

- It is a special type of homogeneous mixture composed of two or more substances.
a. suspension b. **solution** c. colloids d. alloys
- Which of the following is an example of liquid in a solid solution?
a. Air b. Brass c. vinegar d. **dental amalgam**
- You are tasked to analyze a water samples containing contaminants. In what unit of concentration you are going to report the results of analysis?
a. **parts per million** b. percent c. Molarity d. molality
- What is the mole fraction of solvent in 40% isopropyl alcohol solution? (Molar mass isopropyl = 60.1 g/mol; molar mass of water = 18.0 g/mole)
a. 0.16 b. 0.67 c. **0.84** d. 3.3
- What is the concentration, in ppm, if 0.025 g KCl is dissolved in 100 grams of water?
a. 2.5×10^{-4} ppm b. 2.5 ppm c. 4.0×10^2 ppm d. **250 ppm**
- Which of the following statement is TRUE about molality?
a. Molality is temperature dependent
b. Molality is represented by moles / L
c. Molality gives a less accurate value than molarity.
d. **Molality gives a more accurate value than molarity.**
- The term homogeneous mixtures signify that
a. its properties are uniform throughout the mixture.
b. its composition is uniform throughout the mixture.
c. **Both composition and properties are uniform throughout the mixture.**
d. neither composition nor properties are uniform throughout the mixture.
- Which of the following is a quantitative description of the solution?
a. Diluted b. concentrated c. saturated d. **molality**
- Calculate the molarity of liquid HCl will be if density of the solution is 1.17 gm/cc. (MM of HCl= 36.46 g/mol; 1cc = 1 mL, 1 L= 1000 mL)
a. 31.16 M b. **32.09 M** c. 31.16 m d. 32.09 m
- What is the molarity of 20% (w/v) sulfuric acid (H_2SO_4) solution?
a. 0.51 M b. 1.02 M c. **2.04 M** d. 4.08 M
- Which of the following is used as an indicator in the titration of a weak acid and a strong base?
a. Methyl orange (3 to 4.6) c. Methyl red (5 to 6.9)
b. Bromothymol blue (6 to 7.5) d. **Phenolphthalein (8 to 9.6)**
- Arrange the procedure in based on the correct sequence of the titration process.
I. Indicator is added to the solution.
II. A specified amount of analyte is placed on the Erlenmeyer flask.
III. Titrate the sample until a distinct change in the color of the solution is observed.
IV. Titrant is placed in a burette to titrate the sample containing indicator.
a. I, II, III and IV b. II, III, IV and I c. III, IV, II and I d. **II, I, IV and III**
- The following is an indication that the end-point of titration is reached except
a. **The solution becomes dark pink in color.**
b. All the acid has been neutralized by the base.
c. A distinct change in the color of an indicator is noticeable.
d. The solution changes in color from clear to light pink solution.
- What is the concentration of the sulphuric acid solution if 50 ml of 0.5 M $Ba(OH)_2$ solution is added to 100 ml of the solution?
a. **0.25 M** b. 0.5 M c. 25 M d. 50 m
- Find the concentration of HCl, if 10 mL of 0.5 M $Ca(OH)_2$ is required to titrate 50 mL of HCl.
 $2HCl(aq) + Ca(OH)_2(aq) \rightarrow CaCl_2(aq) + 2H_2O(l)$
a. 0.1 M b. **0.2 M** c. 0.5 M d. 1.0 M
- Using the following equation:
 $2NaOH + H_2SO_4 \rightarrow 2H_2O + Na_2SO_4$
How many grams of sodium sulfate will be formed if you start with 1.25 L of a 4.0 M sodium hydroxide? (Atomic mass: Na=23.0, O=16.0, H=1.01, S=32.0.)
a. 1420 g b. 710.0 g c. **155.0 g** d. 0.7040 g

Exam Chapter 2 nuclear Chemistry

Shasha Hu



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Table of Contents Exam Chapter 2nuclear Chemistry

1. Understanding the eBook Exam Chapter 2nuclear Chemistry
 - The Rise of Digital Reading Exam Chapter 2nuclear Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Exam Chapter 2nuclear Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Exam Chapter 2nuclear Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Exam Chapter 2nuclear Chemistry
 - Personalized Recommendations
 - Exam Chapter 2nuclear Chemistry User Reviews and Ratings
 - Exam Chapter 2nuclear Chemistry and Bestseller Lists
5. Accessing Exam Chapter 2nuclear Chemistry Free and Paid eBooks
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- Exam Chapter 2nuclear Chemistry Budget-Friendly Options
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 - Exam Chapter 2nuclear Chemistry Compatibility with Devices
 - Exam Chapter 2nuclear Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Exam Chapter 2nuclear Chemistry
 - Interactive Elements Exam Chapter 2nuclear Chemistry
- 8. Staying Engaged with Exam Chapter 2nuclear Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Exam Chapter 2nuclear Chemistry
- 9. Balancing eBooks and Physical Books Exam Chapter 2nuclear Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Exam Chapter 2nuclear Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Exam Chapter 2nuclear Chemistry
 - Setting Reading Goals Exam Chapter 2nuclear Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Exam Chapter 2nuclear Chemistry
 - Fact-Checking eBook Content of Exam Chapter 2nuclear Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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