

## JEE Advanced 2020 Chemistry| Paper-2 | Code-E

### SECTION 1 (Maximum Marks: 18)

- This section contains **SIX (06)** questions.
- The answer to each question is a **SINGLE DIGIT INTEGER** ranging from **0 TO 9, BOTH INCLUSIVE**.
- For Each question, enter the correct integer corresponding to the answer using the mouse and the on-screen virtual numeric keypad in the place designated to enter the answer.
- Answer to each question will be evaluated according to the following marking scheme:  
*Full Marks* : +3 If **ONLY** the correct numerical value is entered;  
*Zero Marks* : 0 If the question is unanswered;  
*Negative Marks* : -1 In all other cases.

1. The 1<sup>st</sup>, 2<sup>nd</sup>, and the 3<sup>rd</sup> ionization enthalpies,  $I_1$ ,  $I_2$ , and  $I_3$ , of four atoms with atomic numbers  $n$ ,  $n + 1$ ,  $n + 2$ , and  $n + 3$ , where  $n < 10$ , are tabulated below. What is the value of  $n$ ?

Atomic number	Ionization Enthalpy (kJ/mol)		
	$I_1$	$I_2$	$I_3$
$n$	1681	3374	6050
$n + 1$	2081	3952	6122
$n + 2$	496	4562	6910
$n + 3$	738	1451	7733

Answer: 9

Solution:

According to the tabulated data,

Element with Atomic number  $(n + 2)$ , should be an alkali metal

As we see, first ionization enthalpy ( $I_1$ ) is very less but second ionization enthalpy ( $I_2$ ) is very large.

Hence, atomic number can be = 11

That is  $n + 2 = 11$   
 $n = 9$

Note: 'n' can't be '1'.

# Jee Exam Paper 2 Code K Solution

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