

Worksheet 5
General Chemistry II

	Zeroth Order	First Order	Second Order																								
Differential rate law	Rate = $-\frac{\Delta[A]}{\Delta t} = k$	Rate = $-\frac{\Delta[A]}{\Delta t} = k[A]$	Rate = $-\frac{\Delta[A]}{\Delta t} = k[A]^2$																								
Concentration vs. time																											
Integrated rate law	$[A] = [A]_0 - kt$	$[A] = [A]_0 e^{-kt}$ or $\ln[A] = \ln[A]_0 - kt$	$\frac{1}{[A]} = \frac{1}{[A]_0} + kt$																								
Straight-line plot to determine rate constant																											
Relative rate vs. concentration	<table border="1"> <thead> <tr> <th>[A], M</th> <th>Rate, M/s</th> </tr> </thead> <tbody> <tr><td>1</td><td>1</td></tr> <tr><td>2</td><td>1</td></tr> <tr><td>3</td><td>1</td></tr> </tbody> </table>	[A], M	Rate, M/s	1	1	2	1	3	1	<table border="1"> <thead> <tr> <th>[A], M</th> <th>Rate, M/s</th> </tr> </thead> <tbody> <tr><td>1</td><td>1</td></tr> <tr><td>2</td><td>2</td></tr> <tr><td>3</td><td>3</td></tr> </tbody> </table>	[A], M	Rate, M/s	1	1	2	2	3	3	<table border="1"> <thead> <tr> <th>[A], M</th> <th>Rate, M/s</th> </tr> </thead> <tbody> <tr><td>1</td><td>1</td></tr> <tr><td>2</td><td>4</td></tr> <tr><td>3</td><td>9</td></tr> </tbody> </table>	[A], M	Rate, M/s	1	1	2	4	3	9
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Half-life	$t_{1/2} = \frac{[A]_0}{2k}$	$t_{1/2} = \frac{0.693}{k}$	$t_{1/2} = \frac{1}{k[A]_0}$																								
Units of k, rate constant	M/s	1/s	M ⁻¹ s ⁻¹																								

Determining Rate Law:

Series of plots of Concentration v. Time

Determine slope of tangent at t_0 for each plot

Initial Rates

Compare initial rates when [A] changes and [B] is held constant (and vice versa)

Reaction Orders

Substitute initial rates, orders, and concentrations into general rate law rate = $k[A]^m[B]^n$

Rate constant (k) and actual rate law

Temperature Effects

The Arrhenius Equation

$$k = A e^{-E_a/RT}$$

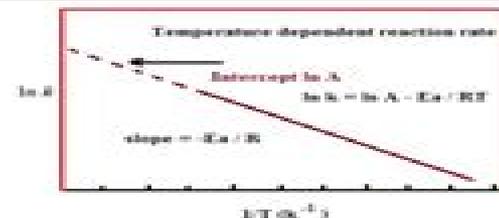
- where k is the rate constant at T
- E_a is the activation energy
- R is the energy gas constant = 8.3145 J/mol K
- T is the Kelvin temperature
- A is the collision frequency factor

$$\ln k = \ln A - E_a/RT$$

$$\ln \frac{k_2}{k_1} = -\frac{E_a}{R} \left(\frac{1}{T_2} - \frac{1}{T_1} \right)$$

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Higher T = larger k = increased reaction rate



General Chemistry Ii Chapter Chemical Kinetics Answers

S Ashworth



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