# General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

## Phase Changes

$\Delta T_b = i K_b m$

$\Delta T_f = i K_f m$

$\pi = i M R T$

$c = k P$

$P_{solution} = x_{solvent} P^o_{solvent}$

$\Delta P = P^o_{solvent} - P_{solution} = x_{solute} P^o_{solvent}$

## Cells

$E = E^o - \frac{RT}{nF} \ln(Q)$

$E = E^o - \frac{0.0592}{n} \log(Q)$ (at 25°C)

$\Delta G = -nF E_{cell}$

$\Delta G^o = -nF E^o_{cell}$

$E^o_{cell} = \frac{RT}{nF} \ln(K)$

$E^o_{cell} = \frac{0.0592 V}{n} \log(K)$ (at 25°C)

### USE ONE:

$E^o_{cell} = E^o_{cathode(red.)} - E^o_{anode(red.)}$

$E^o_{cell} = E^o_{cathode(red.)} + E^o_{anode(red.)}$

## Thermodynamics

$\Delta G = \Delta H - T \Delta S$

$\Delta S_{surr} = \frac{-\Delta H_{sys}}{T}$

$\Delta S_{univ} = \Delta S_{sys} + \Delta S_{surr}$

$\Delta S^o_{rxn} = \Sigma S^o_{products} + \Sigma S^o_{reactants}$

$\Delta G^o_{rxn} = \Sigma \Delta G^o_{products} + \Sigma \Delta G^o_{reactants}$

## Concentrations

$molality = \frac{moles_{solute}}{mass_{solvent}(kg)}$

$\%\ by\ mass = \frac{moles_{solute}}{mass_{solute} + mass_{solvent}} \times 100\%$

## Acids and Bases

$K_w = [H_3O^+][OH^-] = K_a \times K_b$

$pOH = -\log[OH^-]$

$[OH^-] = 10^{-pOH}$

$pH + pOH = pK_a + pK_b = 14.00$ (at 25°C)

$pH = pK_a + \log \frac{[A^-]}{[HA]}$

$pOH = pK_b + \log \frac{[BH^+]}{[B]}$

$\%\ ionization\ \frac{[H_3O^+]_{eq}}{[HA]_0} \times 100\%$

$pK_a = -\log(K_a)$

$pK_b = -\log(K_b)$

$K_a = 10^{-pK_a}$

$K_b = 10^{-pK_b}$

## Equilibria

$Q_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$

$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$

$Q_p = \frac{P_C^c \cdot P_D^d}{P_A^a \cdot P_B^b}$

$K_p = \frac{P_C^c \cdot P_D^d}{P_A^a \cdot P_B^b}$

$K_p = K_c [R \times T]^{\Delta n}$

$\Delta G = \Delta G^o + RT \ln(Q)$

$\Delta G^o = -RT \ln(K)$

$rate = k[A]^x [B]^y$

$k = Ae^{-\frac{E_a}{RT}}$

$\ln(\frac{k_1}{k_2}) = \frac{E_a}{R}(\frac{1}{T_2} - \frac{1}{T_1})$

$half\ life = \frac{\ln(2)}{k}$

$[A]_t = -kt + [A]_o$

$\ln([A]_t) = -kt + \ln([A]_o)$

$\frac{1}{[A]_t} = kt + \frac{1}{[A]_o}$

# General Chemistry Ii Cheat Guide

Michael Rajnik

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