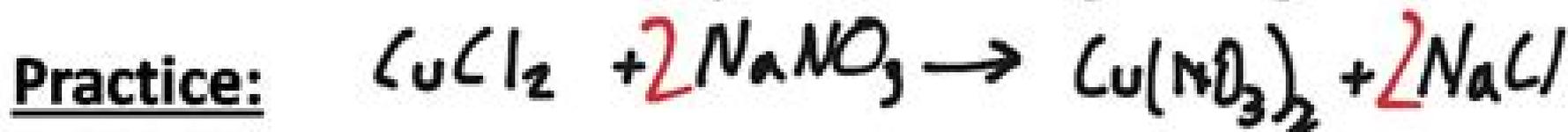


## 9.2 – Stoichiometry – Limiting Reagent



20 grams of copper (II) chloride react with 25 grams of sodium nitrate. Identify the limiting reagent.

$$20\text{g } \cancel{\text{CuCl}_2} \times \frac{1 \text{ mol } \cancel{\text{CuCl}_2}}{137.452\text{g } \cancel{\text{CuCl}_2}} = 0.149 \text{ mol } \cancel{\text{CuCl}_2} \times \frac{2 \text{ mol } \text{NaNO}_3}{1 \text{ mol } \cancel{\text{CuCl}_2}} = 0.298 \text{ mol } \text{NaNO}_3$$

HAVE Need

$$25\text{g } \cancel{\text{NaNO}_3} \times \frac{1 \text{ mol } \cancel{\text{NaNO}_3}}{84.994\text{g } \cancel{\text{NaNO}_3}} = 0.294 \text{ mol } \cancel{\text{NaNO}_3} \times \frac{1 \text{ mol } \text{CuCl}_2}{2 \text{ mol } \cancel{\text{NaNO}_3}} = 0.147 \text{ mol } \text{CuCl}_2$$

You need 0.298 mol  $\text{NaNO}_3$ , but you have 0.294 mol  $\text{NaNO}_3$

# Limiting Reagent Exam Stoichiometry

**Sebastian Brünink**



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