

# General Chemistry

## Chapter 2

### ★ Quantum Mechanics

- Theory using observations as an explanation for something we can NOT see
- Absolutely small matter behaved MUCH differently than a matter of the macroscopic world.
- Subatomic particles appeared to exist in two conditions they had a natural duality
  - EX: How you act individually is different than how you act in a crowd
- Much of the behavior of atoms is determined by the electron
  - Because the electrons are the only subatomic particles NOT in the nucleus

### ★ Nature of Light

- Light is a form of electromagnetic radiation
- All electromagnetic waves move through space at the same constant speed
  - FORMULA on slides
- Amplitude is a measure of light intensity- the larger the amplitude, the brighter the light
- Wavelength is a measure of the distance covered by the wave
- It's 1 point on a wave to the same point on the next wave
- $\lambda$  + Frequency and Wavelengths are inverses of each other as one gets bigger the other gets smaller

### ★ Color

- The color of light is determined by its wavelength or frequency
- White light is a mixture of all the colors of visible light
  - It's a spectrum
  - Red, orange, yellow, green, blue, indigo, violet
- When an object absorbs some of the wavelengths of white light and reflects others, it appears colored; the observed color is predominantly the colors reflected.
- Frequency
  - The number of waves that pass a certain point in a given period of time
  - The number of waves = the number of cycles
- Total Energy-
  - The proportional amplitude of the waves and their frequency
  - The larger the amplitude the more force it has
  - The more frequent the wave strikes, the more force there is
- Long wave - low frequency
- Short wave - high frequency

### ★ Electromagnetic Spectrum

- Visible light is 400 to 700 nm of the electromagnetic spectrum
- Radio wave light has the lowest energy

# General Chemistry Chapter 2nuclear

**Ralph H. Petrucci, Robert K. Wismer**



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