

ACS GENERAL CHEMISTRY FINAL STUDY GUIDE

Avagadro's Number - $N_A = 6.023 \times 10^{23}$

Percent Yield - $(\text{Actual Yield} / \text{Theoretical Yield}) \times 100$

Percent Composition (of an Element) - $[(\text{moles of element} \times \text{molar mass of element}) / (\text{molar mass of compound})] \times 100$

Dilution Equation - $M_1V_1 = M_2V_2$
Mass and Volume

Boyle's Law - $P_1V_1 = P_2V_2$
Pressure and Volume

Charles' Law - $V_1/T_1 = V_2/T_2$
Volume and Temperature

Ideal Gas Law - $PV = nRT$

Volume of an Ideal Gas at STP - 22.4 L

Density (Ideal Gas) - $(P \times \text{Molar Mass}) / (RT)$

Definition of Entahipy - $H = E + PV$

Definition of Heat Capacity - $C = ms$
 $s = \text{specific heat}$

ion - different number of electrons

isotopes - different number of neutrons

Periodic Table Trends - As you go down or left atomic radius increases
As you go up or right ionization energy and electron affinity increases

Ionic Bonds - to charged pieces come together and then there are cations or anions

Covalent Bonds - are when two non metals share electrons

A Compound is probably soluble if it contains these anions: - Cl^-
 Br^-
 I^-

General Chemistry Final Study Guide

Katrin Zwirglmaier



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Table of Contents General Chemistry Final Study Guide

1. Understanding the eBook General Chemistry Final Study Guide
 - The Rise of Digital Reading General Chemistry Final Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Final Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a General Chemistry Final Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Final Study Guide
 - Personalized Recommendations
 - General Chemistry Final Study Guide User Reviews and Ratings
 - General Chemistry Final Study Guide and Bestseller Lists
5. Accessing General Chemistry Final Study Guide Free and Paid eBooks
 - General Chemistry Final Study Guide Public Domain eBooks
 - General Chemistry Final Study Guide eBook Subscription Services
 - General Chemistry Final Study Guide Budget-Friendly Options

6. Navigating General Chemistry Final Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Final Study Guide Compatibility with Devices
 - General Chemistry Final Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Final Study Guide
 - Highlighting and Note-Taking General Chemistry Final Study Guide
 - Interactive Elements General Chemistry Final Study Guide
8. Staying Engaged with General Chemistry Final Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Final Study Guide
9. Balancing eBooks and Physical Books General Chemistry Final Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Final Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry Final Study Guide
 - Setting Reading Goals General Chemistry Final Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry Final Study Guide
 - Fact-Checking eBook Content of General Chemistry Final Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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